

A-LEVEL CHEMISTRY

MARK SCHEME

Practice Paper 1

Maximum marks: 105

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1	(a)	(i)	$[H^+][OH^-]$ OR $[H_3O^+][OH^-]$ <i>Ignore (aq)</i> <i>Must have [] not ()</i>	1
		(ii)	$\sqrt{3.46 \times 10^{-14}}$ ($= 1.86 \times 10^{-7}$) <i>If no square root, CE= 0</i> pH = 6.73 Must be 2dp	2
		(iii)	$[H^+] = 10^{-11.26}$ ($= 4.365 \times 10^{-12}$ OR 4.37×10^{-12}) <i>Mark for working</i> $K_w = [4.365 \times 10^{-12} \text{ OR } 4.37 \times 10^{-12} \times 0.047] = 2.05 \times 10^{-13}$ <i>Allow $2.05 \times 10^{-13} = 2.1 \times 10^{-13}$</i> <i>Mark for answer</i> <i>Ignore units</i>	1
	(b)	(i)	$HCOOH \rightleftharpoons HCOO^- + H^+$ <i>Must have $\rightleftharpoons$ but ignore brackets.</i> OR $HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$ <i>Allow HCO_2^- or $CHOO^-$ ie minus must be on oxygen, so penalise $COOH^-$</i>	1
		(ii)	$K_a = \frac{[H^+][HCOO^-]}{[HCOOH]}$ OR $\frac{[H_3O^+][HCOO^-]}{[HCOOH]}$ <i>Must have all brackets but allow ()</i> <i>Must be $HCOOH$ etc.</i> <i>Allow ecf in formulae from (b)(i)</i>	1
		(iii)	M1 $K_a = \frac{[H^+]^2}{[HCOOH]}$ ($[H^+]^2 = 1.78 \times 10^{-4} \times 0.056 = 9.97 \times 10^{-6}$) <i>Allow HA or HX etc.</i> <i>Allow $[H^+] = \sqrt{K_a \times [HA]}$ for M1</i>	1
			M2 $[H^+] = 3.16 \times 10^{-3}$ <i>Mark for answer</i>	1
			M3 pH = 2.50 allow more than 2 dp but not fewer <i>Allow correct pH from their wrong $[H^+]$ here only if square root shown but not taken, pH = 5.00 can score max 2 for M1 and M3</i>	1

May 200a Level Chemistry Paper

Scott C. Dulebohn



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Table of Contents May 200a Level Chemistry Paper

1. Understanding the eBook May 200a Level Chemistry Paper
 - The Rise of Digital Reading May 200a Level Chemistry Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying May 200a Level Chemistry Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an May 200a Level Chemistry Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from May 200a Level Chemistry Paper
 - Personalized Recommendations
 - May 200a Level Chemistry Paper User Reviews and Ratings
 - May 200a Level Chemistry Paper and Bestseller Lists
5. Accessing May 200a Level Chemistry Paper Free and Paid eBooks
 - May 200a Level Chemistry Paper Public Domain eBooks
 - May 200a Level Chemistry Paper eBook Subscription Services

- May 200a Level Chemistry Paper Budget-Friendly Options
- 6. Navigating May 200a Level Chemistry Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - May 200a Level Chemistry Paper Compatibility with Devices
 - May 200a Level Chemistry Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of May 200a Level Chemistry Paper
 - Highlighting and Note-Taking May 200a Level Chemistry Paper
 - Interactive Elements May 200a Level Chemistry Paper
- 8. Staying Engaged with May 200a Level Chemistry Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers May 200a Level Chemistry Paper
- 9. Balancing eBooks and Physical Books May 200a Level Chemistry Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection May 200a Level Chemistry Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine May 200a Level Chemistry Paper
 - Setting Reading Goals May 200a Level Chemistry Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of May 200a Level Chemistry Paper
 - Fact-Checking eBook Content of May 200a Level Chemistry Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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