

2. Complete the following table by filling in the missing information about isotopes. The first row is completed as an example.

Name of Isotope	Symbol	Mass Number	Number of Protons	Number of Neutrons
hydrogen-3	${}^3_1\text{H}$	3	1	2
scandium-49	${}^{49}_{21}\text{Sc}$	49	21	28
Cobalt - 60	${}^{60}_{27}\text{Co}$	60	27	23
nitrogen-15	${}^{15}_7\text{N}$	15	7	8
Uranium 238	${}^{238}_{92}\text{U}$	238	92	146
Iodine 129	${}^{129}_{53}\text{I}$	129	53	76
Barium-135	${}^{135}_{56}\text{Ba}$	135	56	79
Strontium -86	${}^{86}_{38}\text{Sr}$	86	38	48
Oxygen-18	${}^{18}_8\text{O}$	18	8	10
carbon-14	${}^{14}_6\text{C}$	14	7	7

3. Although oxygen-16 is the most common isotope of oxygen, oxygen-17 and oxygen-18 are also present. Despite the differences in the atomic structures of the three isotopes, there is no difference in how they form ionic or covalent compounds with atoms of other elements. Explain how this can be.

They only differ in the number of neutrons

They have the same electron configurations and only electrons are important for chemical reactions

# Isotopes Practice Set Answer Key

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