

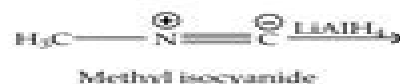
# SOLUTIONS

## Specimen Paper 2023

### Chemistry

#### Section-A

1. (A) (i) anisole, Williamson's synthesis.  
 (ii) three,  $s^{-1}$   
 (iii) lewis acid, lewis base.  
 (iv) acetone, iodoform.  
 (B) (i) Option (b) is correct.  
*Explanation:*



- (ii) Option (c) is correct.

*Explanation:* For a spontaneous reaction, the standard change in standard Gibbs free energy,  $\Delta G^\circ$  and the standard cell potential,  $E_{\text{cell}}^\circ$  will be negative and positive respectively.

- (iii) Option (c) is correct.

*Explanation:* Electronic configuration of  $\text{Cr} = [\text{Ar}] 3d^5 4s^1$  and  $\text{Cu} = [\text{Ar}] 3d^{10} 4s^1$

- (iv) Option (a) is correct.

*Explanation:* Since, for first order reaction, rate law is expressed as

$$k = \frac{2.303}{t} \log \frac{[R_0]}{[R]}$$

At  $t = 8 \text{ min}$ ,

$$k = \frac{2.303}{8} \log \frac{100}{25} = 0.1733 \text{ min}^{-1}$$

For 200 g,

$$k = \frac{2.303}{0.1733} \log \frac{200}{50} = 8 \text{ min}$$

- (C) (i) Phenol  
 (c) Neutral  $\text{FeCl}_3$  solution  
 (ii) Ethylenediamine  
 (d) Bidentate ligand  
 (iii) Colligative property  
 (a) Osmotic pressure  
 (iv) Amino acid  
 (b) Zwitter ion

- (D) (i) Option (a) is correct.

*Explanation:* Conductivity always decreases with the decrease in concentration for both weak and strong electrolytes as the number of ions per unit volume that carry the current

in a solution decreases on dilution.

- (ii) Option (d) is correct.

*Explanation:* Chlorination of nitrobenzene (not the nitration of chlorobenzene) leads to the formation of *m*-nitrochlorobenzene because  $-\text{NO}_2$  group deactivates the ring because it is meta directing.

#### Section-B

2. Osmotic pressure is a colligative property. According to it,

$$\pi = CRT$$

$$\pi = \frac{n}{V} RT \quad \left( C = \frac{n}{V} = \frac{m}{MV} \right)$$

$m$  = Mass of solute = 20 g

$M$  = Molar mass

$V$  = volume of solution = 500 mL

= 0.5 L

$R$  = Gas constant = 0.0821 atm

$\text{mol}^{-1} \text{K}^{-1}$

$T$  = Temperature = 298 K

Place the values in the above equation

$$\pi = \frac{m}{MV} RT$$

$$0.016 = \frac{20 \times 1000}{M \times 500} \times 0.0821 \times 298$$

$$M = \frac{20 \times 1000}{0.016 \times 500} \times 0.0821 \times 298$$

$$= 61164.5 \text{ g}$$

3. (i) Transition elements form a large number of complex compounds due to the comparatively smaller sizes of the metal ions and their high ionic charges and also the availability of

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1. Understanding the eBook Isc Chemistry Paper 24
  - The Rise of Digital Reading Isc Chemistry Paper 24
  - Advantages of eBooks Over Traditional Books
2. Identifying Isc Chemistry Paper 24
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Isc Chemistry Paper 24
  - User-Friendly Interface
4. Exploring eBook Recommendations from Isc Chemistry Paper 24
  - Personalized Recommendations
  - Isc Chemistry Paper 24 User Reviews and Ratings
  - Isc Chemistry Paper 24 and Bestseller Lists
5. Accessing Isc Chemistry Paper 24 Free and Paid eBooks

- Isc Chemistry Paper 24 Public Domain eBooks
- Isc Chemistry Paper 24 eBook Subscription Services
- Isc Chemistry Paper 24 Budget-Friendly Options
- 6. Navigating Isc Chemistry Paper 24 eBook Formats
  - ePub, PDF, MOBI, and More
  - Isc Chemistry Paper 24 Compatibility with Devices
  - Isc Chemistry Paper 24 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Isc Chemistry Paper 24
  - Highlighting and Note-Taking Isc Chemistry Paper 24
  - Interactive Elements Isc Chemistry Paper 24
- 8. Staying Engaged with Isc Chemistry Paper 24
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Isc Chemistry Paper 24
- 9. Balancing eBooks and Physical Books Isc Chemistry Paper 24
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Isc Chemistry Paper 24
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Isc Chemistry Paper 24
  - Setting Reading Goals Isc Chemistry Paper 24
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Isc Chemistry Paper 24
  - Fact-Checking eBook Content of Isc Chemistry Paper 24
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

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- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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