

Experiment 11: Ionic Equilibrium: Solubility and Common Ion Effect

Objectives:

Students should be able to calculate the solubility of a sparingly soluble salt in pure water and in a solution of a common ion. They should also be able to calculate the pH of a solution of a weak acid and a weak base.

Materials: 0.1M NaCl solution

Reagents: 0.1M NaCl solution

Notes:

Table 1: Results	
Initial concentration of NaCl (M)	0.100
Initial concentration of NaCl (M)	0.100
Initial concentration of NaCl (M)	0.100
Initial concentration of NaCl (M)	0.100
Initial concentration of NaCl (M)	0.100

Procedure:

1. Weigh 0.1000 g of NaCl and dissolve it in 100 mL of distilled water in a 100 mL volumetric flask. Make up to the mark with distilled water. $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$

2. Weigh 0.1000 g of NaCl and dissolve it in 100 mL of distilled water in a 100 mL volumetric flask. Make up to the mark with distilled water. $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$

3. Weigh 0.1000 g of NaCl and dissolve it in 100 mL of distilled water in a 100 mL volumetric flask. Make up to the mark with distilled water. $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$

4. Weigh 0.1000 g of NaCl and dissolve it in 100 mL of distilled water in a 100 mL volumetric flask. Make up to the mark with distilled water. $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$

5. Weigh 0.1000 g of NaCl and dissolve it in 100 mL of distilled water in a 100 mL volumetric flask. Make up to the mark with distilled water. $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$

6. Follow the same procedure for the other salts by making the ionic concentration.

7. Some calculations should be done to determine the solubility of the salts in pure water and in a solution of a common ion. The solubility of a salt in a solution of a common ion is lower than in pure water.

Iron Copper Chloride Reaction Lab Answers

J Dewey



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Table of Contents Iron Copper Chloride Reaction Lab Answers

1. Understanding the eBook Iron Copper Chloride Reaction Lab Answers
 - The Rise of Digital Reading Iron Copper Chloride Reaction Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Iron Copper Chloride Reaction Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Iron Copper Chloride Reaction Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Iron Copper Chloride Reaction Lab Answers
 - Personalized Recommendations
 - Iron Copper Chloride Reaction Lab Answers User Reviews and Ratings
 - Iron Copper Chloride Reaction Lab Answers and Bestseller Lists

5. Accessing Iron Copper Chloride Reaction Lab Answers Free and Paid eBooks
 - Iron Copper Chloride Reaction Lab Answers Public Domain eBooks
 - Iron Copper Chloride Reaction Lab Answers eBook Subscription Services
 - Iron Copper Chloride Reaction Lab Answers Budget-Friendly Options
6. Navigating Iron Copper Chloride Reaction Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Iron Copper Chloride Reaction Lab Answers Compatibility with Devices
 - Iron Copper Chloride Reaction Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Iron Copper Chloride Reaction Lab Answers
 - Highlighting and Note-Taking Iron Copper Chloride Reaction Lab Answers
 - Interactive Elements Iron Copper Chloride Reaction Lab Answers
8. Staying Engaged with Iron Copper Chloride Reaction Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Iron Copper Chloride Reaction Lab Answers
9. Balancing eBooks and Physical Books Iron Copper Chloride Reaction Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Iron Copper Chloride Reaction Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Iron Copper Chloride Reaction Lab Answers
 - Setting Reading Goals Iron Copper Chloride Reaction Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Iron Copper Chloride Reaction Lab Answers
 - Fact-Checking eBook Content of Iron Copper Chloride Reaction Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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