

Exercise 1: Density – Preparation of Solutions, KNO₃ and KClO₃

Prepare three separate 250 mL solutions and determine their exact mass

Any substance that is a solid mixture of two or more substances is called a solid solution. Examples include alloys, such as brass and steel, and chemical mixtures, such as cement and concrete. The composition of a solid solution can be expressed in terms of mass percent, volume percent, or mole percent.

1.1. Preparation of solutions

Concentration	Volume of solution (mL)	Mass of solute (g)
10% KNO ₃		
10% KClO ₃		
10% KNO ₃		
10% KClO ₃		

1.2. Preparation of solutions of known concentration

a. Preparation of 10% KNO₃ solution

1. Weigh out 25.0 g of KNO₃ and transfer it to a 250 mL volumetric flask.

2. Add distilled water to the flask until the volume is 250 mL. (The volume of the solution is 250 mL.)

3. The concentration of the solution is 10% KNO₃. (The concentration of the solution is 10% KNO₃.)

4. The concentration of the solution is 10% KNO₃.

5. The concentration of the solution is 10% KNO₃. (The concentration of the solution is 10% KNO₃.)

6. The concentration of the solution is 10% KNO₃.

7. The concentration of the solution is 10% KNO₃.

8. The concentration of the solution is 10% KNO₃.

9. The concentration of the solution is 10% KNO₃.



10. The concentration of the solution is 10% KNO₃.

11. The concentration of the solution is 10% KNO₃.

12. The concentration of the solution is 10% KNO₃.

13. The concentration of the solution is 10% KNO₃.

14. The concentration of the solution is 10% KNO₃.

High School Honors Chemistry Study Guide

E Durkheim



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Table of Contents High School Honors Chemistry Study Guide

1. Understanding the eBook High School Honors Chemistry Study Guide
 - The Rise of Digital Reading High School Honors Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying High School Honors Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High School Honors Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from High School Honors Chemistry Study Guide
 - Personalized Recommendations
 - High School Honors Chemistry Study Guide User Reviews and Ratings
 - High School Honors Chemistry Study Guide and Bestseller Lists
5. Accessing High School Honors Chemistry Study Guide Free and Paid eBooks
 - High School Honors Chemistry Study Guide Public Domain eBooks
 - High School Honors Chemistry Study Guide eBook Subscription Services
 - High School Honors Chemistry Study Guide Budget-Friendly Options
6. Navigating High School Honors Chemistry Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- High School Honors Chemistry Study Guide Compatibility with Devices
- High School Honors Chemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of High School Honors Chemistry Study Guide
 - Highlighting and Note-Taking High School Honors Chemistry Study Guide
 - Interactive Elements High School Honors Chemistry Study Guide
- 8. Staying Engaged with High School Honors Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High School Honors Chemistry Study Guide
- 9. Balancing eBooks and Physical Books High School Honors Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection High School Honors Chemistry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine High School Honors Chemistry Study Guide
 - Setting Reading Goals High School Honors Chemistry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of High School Honors Chemistry Study Guide
 - Fact-Checking eBook Content of High School Honors Chemistry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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