

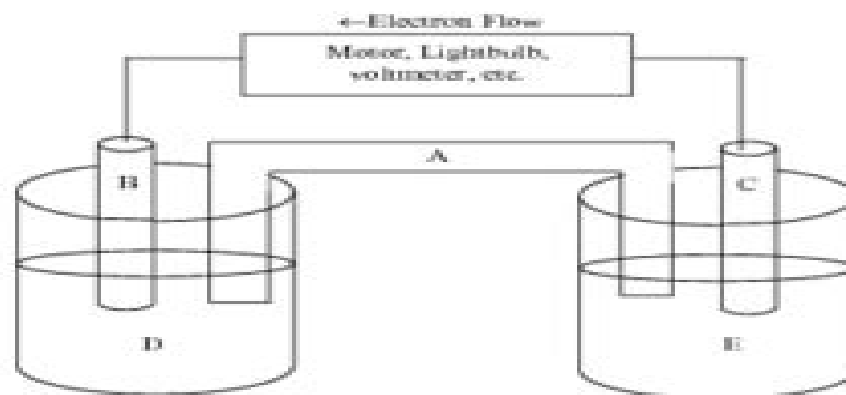
Honors Chemistry

Electrochemistry Problems

Name _____

Date _____ Period _____

The voltaic cell shown above is composed of the metals Nickel (Ni) and Calcium (Ca) and their nitrate salts. Assume both metals are 2^+ in their ionic state. Pay attention to the direction of electron flow. Answer the following questions using the letters indicated above. The same letter may be used more than once.



1. Salt bridge
2. Nickel electrode
3. Calcium nitrate solution
4. Calcium electrode
5. Nickel nitrate solution
6. Anode
7. Cathode
8. Electrode where oxidation occurs
9. Electrode where reduction occurs
10. KNO_3
11. Electrode that gets smaller with time
12. Solution that receives positive ions from the salt bridge
13. If the reduction potentials for Ni and Ca are -0.25V and -2.87V respectively, what is the voltage of the cell?

Write a condensed cell diagram for this reaction.

Write the oxidation half reaction for the cell.

Write the reduction half reaction for the cell.

For the following five questions identify the half reactions as either oxidized or reduced

14. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + 1e^-$
15. $\text{Cu} \rightarrow \text{Cu}^{2+} + 1e^-$
16. $\text{Cl}_2 + 2e^- \rightarrow 2\text{Cl}^-$
17. $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$
18. $2\text{N}^{3-} \rightarrow \text{N}_2 + 6e^-$
19. For the reaction

$$2\text{Mn}^{2+} + 5\text{Cl}_2 \rightarrow 2\text{Mn}^{3+} + 10\text{Cl}^-$$
 which species is the reducing agent? Explain.

High School Electrochemistry Questions

Victor M. Corman



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