

Hewlett-Packard Professional Certification Assessment

Instructions: Fill in each blank with the following problems. Use all appropriate units and proper scientific notation. Use scientific notation for large numbers and negative exponents for small numbers (e.g., 10^{-3}).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_8 + 12\text{O}_2 \rightarrow 10\text{CO}_2 + 8\text{H}_2\text{O}$
 of element then $\text{C}_{10}\text{H}_8 + 12\text{O}_2 \rightarrow 10\text{CO}_2 + 8\text{H}_2\text{O}$ $\Delta H^\circ = -5085 \text{ kJ}$
 and $\text{C}_{10}\text{H}_8 + 12\text{O}_2 \rightarrow 10\text{CO}_2 + 8\text{H}_2\text{O}$ $\Delta H^\circ = -5085 \text{ kJ}$

a. $\Delta H^\circ_{\text{rxn}} = 10\Delta H^\circ_{\text{f,CO}_2} + 8\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_{10}\text{H}_8}$ $\Delta H^\circ_{\text{f,C}_{10}\text{H}_8} = 10\Delta H^\circ_{\text{f,CO}_2} + 8\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

b. $\Delta H^\circ_{\text{rxn}} = 10\Delta H^\circ_{\text{f,CO}_2} + 8\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_{10}\text{H}_8}$ $\Delta H^\circ_{\text{f,C}_{10}\text{H}_8} = 10\Delta H^\circ_{\text{f,CO}_2} + 8\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

2. The standard enthalpy change for the formation of acetylene from its elements is $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$. The standard enthalpy change for the formation of acetylene from its elements is $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$.
 a. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$
 b. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$

a. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$
 b. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$

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 b. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$

c. $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_2} = 226.7 \text{ kJ/mol}$

3. Find the standard enthalpy change for the reaction $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$.
 a. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

Instructions: Find the standard enthalpy change for the reaction $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$.
 a. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

b. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

c. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

d. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

e. $\Delta H^\circ_{\text{rxn}} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{f,C}_2\text{H}_4}$ $\Delta H^\circ_{\text{f,C}_2\text{H}_4} = 2\Delta H^\circ_{\text{f,CO}_2} + 2\Delta H^\circ_{\text{f,H}_2\text{O}} - \Delta H^\circ_{\text{rxn}}$

Hess Law Problems With Answers

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