

UNIT 1

11.1

- Condensed phases- a phase in which the particles are packed in close proximity to one another
- IMFs: solid>liquid>gas
- IMFs are the forces that exist between the molecules, atoms, or ions that make up a particular substance
- Stronger the IMFS, that harder it is to separate the particles.
- Fusion: solid $\rightarrow$ liquid and Vaporization: liquid $\rightarrow$ gas both require an input of energy
- Condensation: gas $\rightarrow$ liquid releases energy
- Enthalpy of vaporization ΔH_{vap} is the energy required to convert one mole of a liquid to vapor at a given temperature
- As IMF $\uparrow = \Delta H_{\text{vap}} \uparrow$



11.2

- Vapor Pressure- the pressure of the vapor of a substance in contact with its liquid or solid phase in a sealed container
 - When liquid is placed in a sealed container, some molecules escape into vapor phase (increasing vapor pressure) and some will recondense into liquid phase. When these become equal, they reach the dynamic equilibrium
 - Equilibrium vapor pressure is the pressure of vapor of a substance at equilibrium in contact with its liquid or solid phase in a sealed container
- $\uparrow$ temperature = $\uparrow$ VP
- $\uparrow$ IMFs = $\downarrow$ VP
- Boiling Point of a liquid is when the Vapor pressure of the liquid is equal to the external pressure
 - Normal boiling point is the temperature when the vapor pressure of a liquid is equal to 1 atm (760 mmHg)
- $\uparrow$ IMFs = $\uparrow$ BP

$$\ln P = \frac{-\Delta H_{\text{vap}}}{RT} + C \quad \ln \frac{P_2}{P_1} = \frac{-\Delta H_{\text{vap}}}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$$

- $R = 8.314 \times 10^{-3} \text{ kJ/K} \cdot \text{mol}$, $C = \text{constant of integration}$
 - $\ln P$ vs. $1/T$ and the slope is $-\Delta H_{\text{vap}}/R$

11.3

- Cohesive forces: a force of attraction between molecules of a single substance and affect the physical properties of a liquid
- Surface tension is the measure of force required to "break" the surface of a liquid
 - Results from different IMFs experienced by molecules at the surface of a liquid and those in the interior of the liquid

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