

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?



- Cl^-
- SO_4^{2-}
- NH_4^+
- HCO_3^-
- H_3O^+

2. In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?

17

- H_2O and HCN
- H_2O and OH^-
- CN^- and H_2O
- HCN and OH^-

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

17

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

17

- | | |
|----|------|
| a. | 1.5 |
| b. | 2.5 |
| c. | 3.5 |
| d. | 12.5 |

5. The pH of a 1.0×10^{-3} M $\text{Ba}(\text{OH})_2$ solution at 25°C is

17

- | | |
|----|------|
| a. | 2.7 |
| b. | 3.0 |
| c. | 11.0 |
| d. | 11.3 |

6. Which is the strongest acid?



- HClO
- HClO_2
- HClO_3
- HClO_4

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

17

- a. $\text{NaC}_2\text{H}_3\text{O}_2$
b. NaNO_3
c. NH_4Cl
d. BaSO_4

8. The weakest of the bases listed is

Acid	Conjugate Base	K _a (Ionization Constant of Acid)
HCl	Cl ⁻	100% ionized
HSO ₄ ⁻	SO ₄ ²⁻	1.2 x 10 ⁻²
H ₂ S	HS ⁻	5.7 x 10 ⁻⁸
HS ⁻	S ²⁻	1.2 x 10 ⁻¹⁴



- Cl^-
- CN^-
- HS^-
- S^{2-}
- SO_3^{2-}

Figure 1: Schematic representation of the experimental design. The diagram shows a sequence of events: 'Stimulus' (a horizontal bar with a central dot), 'Response' (a horizontal bar with a central dot), and 'Feedback' (a horizontal bar with a central dot). The 'Stimulus' bar is labeled 'Stimulus' and 'Response' and 'Feedback' are labeled 'Response' and 'Feedback' respectively. The 'Stimulus' bar is divided into segments labeled 'Stimulus' and 'Response' and 'Feedback' respectively. The 'Response' bar is divided into segments labeled 'Response' and 'Feedback' respectively. The 'Feedback' bar is divided into segments labeled 'Response' and 'Feedback' respectively.

1000

100

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

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- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

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- a 0.100 M sodium solution,
- a 0.0500 M sodium hydroxide solution,
- a 0.0500 M benzoic acid solution,
- a 0.0500 M sodium benzoate solution.

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Lingjun Ying



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