

Equilibrium Constants in Terms of Pressure, K_p

When the reactants and products are gases, the equilibrium constant can be expressed in terms of partial pressures (K_p).



The relationship between K_p and K_c is given by the following equation:

$$K_p = K_c(RT)^{\Delta n}$$

$$\Delta n = \sum (\text{product coefficients}) - \sum (\text{reactant coefficients})$$

Example: $3\text{H}_2(g) + \text{N}_2(g) \rightleftharpoons 2\text{NH}_3(g) \quad K = 5.80 \text{ at } 180^\circ\text{C} \quad K_p = ?$

$$\Delta n = 2 - (3 + 1) = -2 \quad K_p = 5.80 \times (0.08206 \times 273)^{-2} = 4.54 \times 10^{-3}$$

- If the number of gas molecules does not change in the reaction, $K_p = K_c$ because $\Delta n = 0$.

K for Heterogeneous Equilibria

The concentrations of solids, liquids, and the solvent are omitted from equilibrium expressions.



Equilibrium Constants and Units

The more accurate expression for the equilibrium constant is given by using the activities of reactants and products.

$$K = \frac{(a_c)^c (a_d)^d}{(a_a)^a (a_b)^b}$$

For most undergraduate general chemistry courses, it is assumed that $a \approx 1$ since the differences are small and can be neglected. Because activities have no units, K is traditionally shown without units.

The Direction of Reaction and Quotient (Q)

To predict the direction of reaction before an equilibrium is reached, reaction quotient is used. It is similar to the expression for K , however, the **actual concentrations are used**.



- $Q < K$ Reaction has a tendency to form more products.
- $Q > K$ Reaction has a tendency to form more reactants.
- $Q = K$ Reaction is at equilibrium.

$$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b} \quad Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

- If any reactant or product is missing in the initial mixture, the reaction will shift in the direction forming some of that component.

Le Châtelier's Principle

If a chemical system at equilibrium is disturbed, the system shifts in a direction that minimizes the disturbance. This can be done in different ways.

Changing Concentrations at Equilibrium

- Adding a reactant produces more product by consuming some of the reactant.
- Adding a product produces more reactant by consuming some of the product.



Changing The Reaction Volume

- Decreasing the volume shifts the reaction to in the direction with less gas molecules.
- Increasing the volume shifts the reaction to in the direction with more gas molecules.



Changing The Pressure

- The pressure can be changed in a few ways: 1) By decreasing the volume (see above), 2) By adding more volume to the mixture (see above), 3) By adding an inert gas.
- Adding an inert gas does not cause a shift in equilibrium because it does not alter the partial pressures of any of the reacting components.

Changing The Temperature

Know the sign of ΔH is different than what we write for heat!

For an endothermic reaction, treat the heat as a product.



For an exothermic reaction, treat the heat as a reactant.



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Judd E. Hollander



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