

Acid/Base Review Sheet
Chemistry I

Name: Key 2012

1. What is the difference between the Arrhenius definition and the Bronsted-Lowry definition of an acid and a base?

- Arrhenius definition has the acid/base in solution.
Acids increase $[H^+]$ and bases increase $[OH^-]$
- B L acids are proton donors, bases are proton acceptors

2. Identify the acid, base, conjugate acid and conjugate base in the following:



3. Define the following terms:

Arrhenius Acid - substance that increases $[H^+]$ in solution

Arrhenius Base - substance that increases $[OH^-]$ in solution

Bronsted-Lowry Acid - proton donor
(H^+)

Bronsted-Lowry Base - proton acceptor
(H^+)

Monoprotic acid - an acid with one proton ex: HCl , HNO_3

Diprotic acid - an acid with two protons ex: H_2SO_4

General Chemistry Acid Base Review Sheet

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