ssessment #3 Review	Name:
ou must show all work for	credit.
ue Date:	(No late packets will be accepted)
I. Balance Equation	ans.
Balance the equations	
 Mg + ∠ H₂O → M; 	$g(OH)_2 + H_2$
 Mg + ZNaBr → N 	MgBr ₂ + ZNa
3. C ₃ H ₈ + 5 O ₂ → 3 C	
4. 4 P + 5 O ₂ → 2 P ₂ O ₅	1 1 K CO 1 7 EVOUN
5. Fe ₂ (SO ₄) ₃ + 6 KOH	I →3 K₂SO₄ + Z Fe(OH)₃
 Fe₂O₃ + 3 H₂ → F 	7
7 How many moles of	hydrogen gas react with 1 mole of Fe ₂ O ₃ in reaction #6?
7. 11011 1112117	
rite the balanced equation	n for problems # 8 and 9:
Hydrogen gas and n	nitrogen monoxide combine to form water and nitrogen gas
O Mars and makes	phate and calcium chloride are mixed they form calcium phosphate and
sodium chloride	2 Nas Poy 13 Cacle - Cas (Poy) = + & Na C
10. Using the following b	balanced equation, prove (or explain) the Law of Conservation of Mass.
$N_2 + 3H_2 \rightarrow 2NH_3$	
	2 N on lett 2 on ight
	su no 1) gods created or lost
44 Union the belonged	equation shown in problem #10, explain what the coefficients represent.
11. Using the balanced	Al Male s
12. Using the balanced	equation shown in problem #10, list the products of the reaction.
	r)H.T
42 Heine the belonged	counting shown in problem #10. list the reactants of the reaction
13. Using the balanced	equation shown in problem #10, list the reactants of the reaction.
13. Using the balanced	
II. Types of Reacti	ions
II. Types of Reacti	ions e of reaction.
II. Types of Reacti	ions e of reaction.
II. Types of Reacti	ions e of reaction.
II. Types of Reacti irections: Identify the type = synthesis = decomposition	ions e of reaction. SD = single displacement C= combustion DD = double displacement
II. Types of Reacti irections: Identify the type = synthesis = decomposition	ions e of reaction.
II. Types of Reacti irections: Identify the type = synthesis = decomposition 14. HCI + MgS	ions e of reaction. SD = single displacement C= combustion DD = double displacement
II. Types of Reacti irections: Identify the type = synthesis = decomposition 14. ☐ HCI + MgS 15. ☐ Cu₂(CO₃)₃ →	ions e of reaction. SD = single displacement DD = double displacement C= combustion C= combustion C= combustion C= combustion
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II. Types of Reacting the type is a synthesis of a decomposition 14. HCl + MgS 15. Cu₂(CO₃)₃ → 16. C₂H₅ + O₂ →	ions e of reaction. SD = single displacement DD = double displacement SO ₄ → MgCl ₂ + H ₂ SO ₄ Cu + CO ₂ + O ₂ → CO ₂ + H ₂ O P ₂ O ₅
II. Types of Reactions: Identify the type S = synthesis D = decomposition 14.	ions e of reaction. SD = single displacement DD = double displacement $C = combustion$ $C $

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