

Can You Find the Side Sides?

In this experiment, boiling acids (H₂SO₄/H₂O) reacts with hydrocarbons and (HCl) to produce various alkenes, various alcohols, and water. You will determine the moles of reactants used and moles of product produced using both the lab data and stoichiometry and compare the two.

Sample Table:

moles of reactant used and boiling acids	moles
moles of isopropyl alcohol used	moles
moles of isopropyl alcohol	moles
moles of 1-propanol (or isopropanol) produced	moles
moles of isopropyl alcohol used	moles
moles of isopropanol produced	moles

Calculations Questions:

- Calculate the number of moles of boiling acids used in this lab. (10%)

$$\text{moles of } \text{H}_2\text{SO}_4 = \frac{\text{mass of } \text{H}_2\text{SO}_4}{\text{molar mass of } \text{H}_2\text{SO}_4}$$
- Calculate the number of moles of product alcohols produced in this lab. (10%)

$$\text{moles of } \text{C}_3\text{H}_8\text{O} = \frac{\text{mass of } \text{C}_3\text{H}_8\text{O}}{\text{molar mass of } \text{C}_3\text{H}_8\text{O}}$$
- What is the experimental molar ratio of boiling acids (H₂SO₄/H₂O) to product alcohols (C₃H₈O)? (10%)

$$\text{molar ratio} = \frac{\text{moles of } \text{H}_2\text{SO}_4/\text{H}_2\text{O}}{\text{moles of } \text{C}_3\text{H}_8\text{O}}$$
- What is balanced equation for the reaction that took place in this experiment? (10%)

$$\text{isopropanol} + \text{H}_2\text{SO}_4 \rightarrow \text{isopropanol} + \text{H}_2\text{O}$$
- According to the balanced equation, what is the theoretical molar ratio of boiling acids to product alcohols? (10%)

$$1:1$$
- What the experimental molar ratio used by this reaction for theoretical molar ratio? If not, give some possible reasons. (10%)

$$\text{theoretical molar ratio} = \frac{\text{moles of } \text{H}_2\text{SO}_4/\text{H}_2\text{O}}{\text{moles of } \text{C}_3\text{H}_8\text{O}}$$

Episode 801 Chemistry Answers

M Lipman



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