

Sulfur clock lab

Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1.0M	0.1M H_2O_2 , 0.01M KSCN	Reaction setup
B	0.5M	0.1M H_2O_2 , 0.01M KSCN	0.1M, 0.01M, 0.01M
C	0.25M	0.1M H_2O_2 , 0.01M KSCN	0.1M, 0.01M, 0.01M
D	0.125M	0.1M H_2O_2 , 0.01M KSCN	0.1M, 0.01M, 0.01M

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the calculated amount of H_2O_2 solution to the reaction mixture using a graduated cylinder.
6. Add exactly 1.00 mL of 0.1M KSCN solution to the reaction mixture.
7. Prepare and record the data for the reaction.
8. Prepare solutions A, B, C, D and E as shown:
 - a. Transfer 10.0 mL of 0.1M H_2O_2 solution to the reaction mixture
 - b. Add 1.0 mL of 0.1M KSCN solution to the reaction mixture
 - c. Dilute to the 100 mL mark using distilled water
 - d. Stop the reaction and record the time taken
9. Transfer solution A to the reaction mixture as 100 mL of H_2O_2 solution. Do not stop the reaction and record the time.

Lab Sulfur Clock Answers

Terry C. Jones

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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