

## Lab Report Experiment 11: Chemical Kinetics

**Data Table I (Rate vs Concentration)** The temperature is 45 °C.

Trial	H <sub>2</sub> O (mL)	0.010 M H <sub>2</sub> O <sub>2</sub> (mL)	1.0 M HCl (mL)	0.050 M KI (mL)	Time (sec)	Rate (1/sec)
1	60.0	10.0	20.0	10.0	45	
2	50.0	20.0	20.0	10.0	23	
3	40.0	10.0	40.0	10.0	45	
4	50.0	10.0	20.0	20.0	24	

**Date Table II (Rate vs Temperature)**

Trial	H <sub>2</sub> O (mL)	0.010 M H <sub>2</sub> O <sub>2</sub> (mL)	1.0 M HCl (mL)	0.050 M KI (mL)	Temp °C	Time (sec)	Rate (1/sec)
1	60.0	10.0	20.0	10.0	5	67	
2	60.0	10.0	20.0	10.0	15	60	
3	60.0	10.0	20.0	10.0	25	55	
4	60.0	10.0	20.0	10.0	35	52	
5	60.0	10.0	20.0	10.0	45	45	

### Analysis:

1. What happened to the rate when the amount of H<sub>2</sub>O<sub>2</sub> was doubled? What is the order of the reaction with respect to H<sub>2</sub>O<sub>2</sub>?
2. What happened to the rate when the amount of HCl was doubled? What is the order of the reaction with respect to HCl?

# Kinetics Of A Reaction Lab Report Answers

**M Woodhall**



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