

NAME: _____

CLASS: _____

DATE: _____

1. Balance the following reaction:
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$

- | | | | |
|-------------------------|---------|-------------------------|---------|
| ☐ A | 1,2,1,2 | ☐ B | 1,2,2,2 |
| ☐ C | 2,1,1,2 | ☐ D | 1,2,1,1 |

2. Using the balanced equation in question 1,
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
 How many grams of HCl are consumed by the reaction of 2.50 moles of magnesium?

- | | | | |
|-------------------------|--------|-------------------------|-------|
| ☐ A | 218 g | ☐ B | 7.29 |
| ☐ C | 5.00 g | ☐ D | 182 g |

3. Using the balanced equation in question 1,
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
 What is the mass in grams of H_2 gas when 4.0 moles of HCl is added to the reaction?

- | | | | |
|-------------------------|--------|-------------------------|-------|
| ☐ A | 0.99 g | ☐ B | 4.0 g |
| ☐ C | 2.0 g | ☐ D | 6.0 g |

4. Using the balanced equation in question 1,
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$
 How many moles of HCl are consumed in the production of 7.5 moles of MgCl_2 ?

- | | | | |
|-------------------------|----------|-------------------------|-----------|
| ☐ A | 15 moles | ☐ B | 3.8 moles |
| ☐ C | 23 moles | ☐ D | 7.5 moles |

5. Balance the following reaction :
 $\text{CaC}_2\text{(s)} + \text{H}_2\text{O(l)} \rightarrow \text{C}_2\text{H}_2\text{(g)} + \text{Ca(OH)}_2\text{(aq)}$

- | | | | |
|-------------------------|---------|-------------------------|---------|
| ☐ A | 2,1,1,1 | ☐ B | 2,1,2,1 |
| ☐ C | 1,2,1,1 | ☐ D | 1,2,2,2 |

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