

## Topic 2: Chemical Change (P 114-129)

## Unit 1: Physical &amp; Chemical Change

- Physical change = change in matter that can be reversed.
- 2 substances that are mixed and can be separated by physical means.
- E.g. solid water (ice) is arranged in crystal lattice.
  - Heat breaks the lattice - transforming it into liquid water and gas (water vapor).
  - Atoms in water remain bonded due to strong forces.
- Chemical change = changing something so that it can't be changed back to its original state.
- E.g. heating iron and sulphur together creates a brown, sticky substance called iron sulphide (FeS).
  - This new substance has different properties from the old one (FeS isn't magnetic).
  - Heating sulphur with iron causes atoms to break apart and rejoin in a strong chemical bond - new lattice formed.



- Chemical changes require more energy than physical ones.

| Physical Change                  | Chemical Change                  |
|----------------------------------|----------------------------------|
| Easy to reverse.                 | Difficult to reverse.            |
| No new chemical substances form. | New chemical substances form.    |
| Require small amounts of energy. | Require large amounts of energy. |

## Unit 2: Conservation of Atoms During Chemical Change

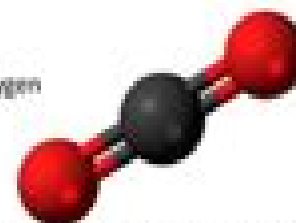
## Law of Conservation of Matter

- No. of atoms doesn't change when a chemical reaction occurs.
- No. molecules may change.
- Law of conservation of matter = mass of reactants and mass of products stays the same.

## Law of Constant Proportions

- Law of constant proportions = a chemical compound always contains the same elements in fixed proportions by mass.

- E.g. CO<sub>2</sub>:
  - Relative atomic mass of carbon = 12.
  - Relative atomic mass of oxygen = 16 (2 oxygen atoms required in CO<sub>2</sub>).
  - Total relative atomic mass of CO<sub>2</sub> is 12 + 16 + 16 = 44.
  - Ratio of carbon to oxygen remains constant - any mass of CO<sub>2</sub> will be  $\frac{12}{44}$  carbon and  $\frac{32}{44}$  oxygen.
  - Ratio of mass of carbon to oxygen is always 12:32 or 3:8 in any amount of CO<sub>2</sub>, irrespective of how it is formed.



## Unit 3: Representing Chemical Change

## Balanced Chemical Equations

- Chemical change is represented by chemical equations.
- Substances on the left of the arrow = reactants.
- Substances on the right of the arrow = products.

# Grade 10 Physical Science 2014 Exemplar

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## **Table of Contents Grade 10 Physical Science 2014 Exemplar**

1. Understanding the eBook Grade 10 Physical Science 2014 Exemplar
  - The Rise of Digital Reading Grade 10 Physical Science 2014 Exemplar
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade 10 Physical Science 2014 Exemplar
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade 10 Physical Science 2014 Exemplar
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 10 Physical Science 2014 Exemplar
  - Personalized Recommendations
  - Grade 10 Physical Science 2014 Exemplar User Reviews and Ratings
  - Grade 10 Physical Science 2014 Exemplar and Bestseller Lists
5. Accessing Grade 10 Physical Science 2014 Exemplar Free and Paid eBooks

- Grade 10 Physical Science 2014 Exemplar Public Domain eBooks
  - Grade 10 Physical Science 2014 Exemplar eBook Subscription Services
  - Grade 10 Physical Science 2014 Exemplar Budget-Friendly Options
6. Navigating Grade 10 Physical Science 2014 Exemplar eBook Formats
- ePub, PDF, MOBI, and More
  - Grade 10 Physical Science 2014 Exemplar Compatibility with Devices
  - Grade 10 Physical Science 2014 Exemplar Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Grade 10 Physical Science 2014 Exemplar
  - Highlighting and Note-Taking Grade 10 Physical Science 2014 Exemplar
  - Interactive Elements Grade 10 Physical Science 2014 Exemplar
8. Staying Engaged with Grade 10 Physical Science 2014 Exemplar
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade 10 Physical Science 2014 Exemplar
9. Balancing eBooks and Physical Books Grade 10 Physical Science 2014 Exemplar
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade 10 Physical Science 2014 Exemplar
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Grade 10 Physical Science 2014 Exemplar
- Setting Reading Goals Grade 10 Physical Science 2014 Exemplar
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 10 Physical Science 2014 Exemplar
- Fact-Checking eBook Content of Grade 10 Physical Science 2014 Exemplar
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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