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		н	ш	ш	

An atom is a really tiny particle. It has a radius of about 0.1 nanometres (1x10-10 m)

Atoms are neutral because the no. protons

= no. electrons An atom consists of:

Particle	Relative	Charg
Proton	1	+1
Neutron	1	0
Electron	0 (very	-1

Tour can find the

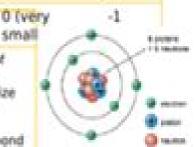
of Mumber of meutrons - ato-

THE NUCLEUS

- •Middle of the atom
- Contains protons and neutrons
- Has a radius of around 1x10-14m (1/100° of atom radius)
- Positively charged due to protons
- Almost whole mass of atom concentrated in mucleus.

THE ELECTRONS

- Move around nucleus in electron shells
- -Negatively charged & tiny, but cover a lot of space
- •The volume of their orbits determines the size of the atoms
- Negatively charged
- ·Bond to other electrons to make chemical bond



Element

A substance made up of atoms that all have the same number of protons in their nucleus (i.e. all are the same atom)

Different forms of the same Isotopes element, which have the same number of protons but a different number of neutrons.

2 common examp -Carbon-12 *Carbon-13

Because many

elements can exist as a number of different

isotopes... relative atomic mass (A) is

Carbon-13 Carlotte. and Property. ALL PROPERTY. # s. Streetman in Thireman Machine anadom Market breaker 18.16 44.47 - 10 + 111

To balance an equation, you should write out all the elements in the middle, write out how many there are on either side and add numbers to balance it

CHEMISTRY CT

Atomic structure, separating

H₂SO₄ + NaOH Na/50++ H₂O

BQ) Copper has two stable isotopes. Cu-63 has an abundance of 69.2% and Cu-65 has an abundance of 30.8%. Calculate the relative atomic mass of copper to 1dp.

(882-1)-(383-65)_ 61584-2981_43614 63.616 = 63.6

> The relative atomic mass is an average mass taking into account the different masses & abundances (amounts) of all the isotopes that make up the

to enable

Compound

Substances formed from 2/more elements, where the atoms of each are held together by chemical bonds.

- Making bond involves atoms giving away, taking or sharing electrons
- A chemical bond is required to break these bonds and separate the compound
- There are 2 types of compound bonds:

1) IONIC BONDING

- · When a metal and a non-metal react to form compound
- The metal loses electrons to form a positive ion
- The non-metal gains electrons to form a negative ion
- The opposite charges of the ions. means they're strongly attracted

2) COVALENT BONDING

- A compound formed from nonmetal consists of molecules
- Each atoms shares electrons with another atom

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used instead of mass number when referring to the element

relative alonic mass = sum of (instage abundance + isotope mass number)

sum of all isotope abundances

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