

Definitions

Atom -

Particles that make up all substances.

Element -

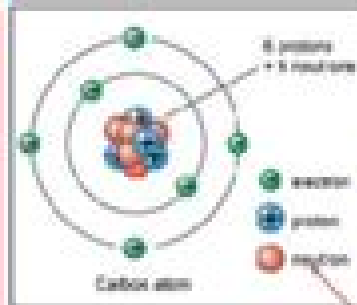
A substance made up of only one kind of atom.

Compound -

A substance made of different types of atoms joined together.

Mixture -

A substance consisting of atoms that can be easily separated as they are not joined together.



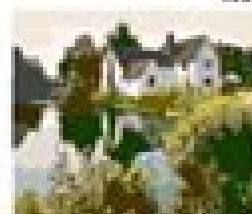
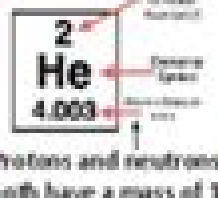
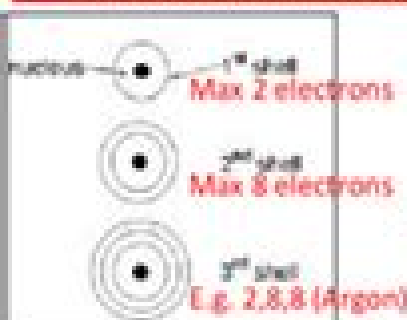
Periodic Table of the Elements

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72
73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90
91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108

No charge

Group 0 or 8 are **UNREACTIVE**
They have a full outer shell of electrons

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)



Used for building

LIMESTONE

Calcium Carbonate CaCO_3

HEAT

(thermal decomposition)

CaO

Calcium Oxide

Limewater

The test for CO_2

Ca(OH)_2

Or Calcium Hydroxide

comes from adding water to CaO

Carbonates
Decompose when
heated
= Oxide + CO_2
React with acid
= Salt + water + CO_2

Balanced Equations



C=1

H=4

O=4

C=1

H=4

O=4

A balanced equation has the same number of atoms of each element on both sides

Using Molecules: In this equation, the large number in front of the chemical symbol tells you the molecules needed.

e.g. 1 CH_4 molecule reacts with 2 O_2 molecules

Covalent Bonding

Non-metals sharing electrons



Ionic Bonding



IONS
Charged particles

Good: More jobs
Bad: Destroy habitats and Landscape
Would you want to live next to one?

TRANSITION METALS



Metals - Very useful
e.g. Copper wires conduct well
Most need to be **ALLOYED**
to make them harder

Metals which are un-reactive are found in their **NATIVE STATE** e.g. **GOLD**
More reactive metals are found as **METALS ORES** and need to be **EXTRACTED**



C1 CHEMISTRY

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Jane Morris



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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Chemistry C1 March 2013 Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Chemistry C1 March 2013 Past Papers

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 - Gcse Chemistry C1 March 2013 Past Papers User Reviews and Ratings
 - Gcse Chemistry C1 March 2013 Past Papers and Bestseller Lists
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 - Interactive Elements Gcse Chemistry C1 March 2013 Past Papers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Chemistry C1 March 2013 Past Papers
 9. Balancing eBooks and Physical Books Gcse Chemistry C1 March 2013 Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Chemistry C1 March 2013 Past Papers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gcse Chemistry C1 March 2013 Past Papers
 - Setting Reading Goals Gcse Chemistry C1 March 2013 Past Papers
 - Carving Out Dedicated Reading Time
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- Fact-Checking eBook Content of Gcse Chemistry C1 March 2013 Past Papers
- Distinguishing Credible Sources

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- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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