

TITRATIONS



Calculate the concentration of the sulfuric acid in mol/dm³.

Use only the student's concordant results.

$$\text{mean titre} = \frac{22.10 + 22.15 + 22.15}{3} = 22.13 \text{ cm}^3 \quad \checkmark$$

$$n(\text{NaOH}) = 0.105 \times (22.13/1000) = 0.002324 \text{ mol} \quad \checkmark$$

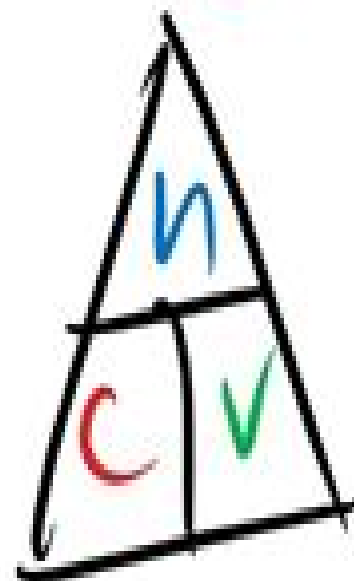
$$n(\text{H}_2\text{SO}_4) = 0.002324 / 2 = 0.001162 \text{ mol} \quad \checkmark$$

$$c(\text{H}_2\text{SO}_4) = \frac{0.001162}{(25/1000)} = 0.0465 \text{ mol/dm}^3 \quad \checkmark$$

$$n = m/M_r$$

$$n = C \times V$$

$$C = n/V$$



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R Bogdan

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