

Experiment 23**Advance Study Assignment:** Determination of the Equilibrium Constant for a Chemical Reaction

1. A student mixes 5.00 mL 2.00×10^{-3} M $\text{Fe}(\text{NO}_3)_3$ in 1 M HNO_3 with 3.00 mL 2.00×10^{-3} M KSCN and 2.00 mL of water. She finds that in the equilibrium mixture the concentration of FeSCN^{2+} is 7.0×10^{-5} M. Find K_c for the reaction $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$.

Step 1 Find the number of moles Fe^{3+} and SCN^{-} initially present. (Use Eq. 3.)

Fe^{3+}

_____ moles Fe^{3+} ; _____ moles SCN^{-}

Step 2 How many moles of FeSCN^{2+} are in the mixture at equilibrium? What is the volume of the equilibrium mixture? (Use Eq. 3.)

_____ mL; _____ moles FeSCN^{2+}

How many moles of Fe^{3+} and SCN^{-} are used up in making the FeSCN^{2+} ?

_____ moles Fe^{3+} ; _____ moles SCN^{-}

Step 3 How many moles of Fe^{3+} and SCN^{-} remain in the solution at equilibrium? (Use Eq. 4 and the results of Steps 1 and 2.)

_____ moles Fe^{3+} ; _____ moles SCN^{-}

Step 4 What are the concentrations of Fe^{3+} , SCN^{-} , and FeSCN^{2+} at equilibrium? (Use Eq. 3 and the results of Step 2 and Step 3.)

$[\text{Fe}^{3+}] =$ _____ M; $[\text{SCN}^{-}] =$ _____ M; $[\text{FeSCN}^{2+}] =$ _____ M

Step 5 What is the value of K_c for the reaction? (Use Eq. 2 and the results of Step 4.)

$K_c =$ _____

(continued on following page)

Experiment 23 Determination Equilibrium Constant Answers

VM Jensen



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