

Experiment 25

Observations and Calculations: pH: Buffers and Their Properties

A. Determination of pH by the Use of Acid-Base Indicators

Indicator	Color with 0.1 M Solution of			
	HCl	NaH_2PO_4	$\text{HC}_2\text{H}_3\text{O}_2$	ZnSO_4
Methyl violet	blue	dark violet	dark purple	purple
Thymol blue	red	yellow	orange	yellow
Methyl yellow	red	yellow	red-orange	yellow
Congo red	purple	brown	dark purple	red-orange
Bromocresol green	yellow	green	yellow	green
pH range	1-3	2-4	2-3	2-5

Circle the observation(s) for each solution that was (were) most useful in estimating the pH range.

B. Measurement of the pH of Some Typical Solutions

Record the pH and the color observed with bromocresol green for each of the 0.1 M solutions that were tested.

	NaCl^1	Na_2CO_3^2	$\text{NaC}_2\text{H}_3\text{O}_2^3$	NaHSO_4^4
pH	4.94	10.97	7.17	1.67
Color	green	blue	blue	yellow

For any two solutions having a pH less than 6 or greater than 8, write a net ionic equation to explain qualitatively why the solution has that pH.

Solution Na_2CO_3 Equation $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$

Solution NaHSO_4 Equation $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons (\text{SO}_4)^{2-} + \text{H}_3\text{O}^+$

Explain why the color observed with bromocresol green for each of the four solutions is reasonable, given the pH.

C. Some Properties of Buffers

Buffer system selected $\text{HCO}_3^- - \text{CO}_3^{2-}$ HB is hydrogen carbonate-carbonate (name the acid)

1. pH of buffer 9.28 $[\text{H}^+] 5.24 \times 10^{-10}$ M pK_a (by Eq. 4b) _____

2. pH of diluted buffer 9.24 $[\text{H}^+] 5.75 \times 10^{-10}$ M pK_a _____

pH after addition of 5 drops NaOH 9.34

pH after addition of 5 drops HCl 9.25

Comment on your observations in Parts 1 and 2.

(continued on following page)

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SJ Ball



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Table of Contents Experiment 25ph Measurements Buffers And Their Properties

1. Understanding the eBook Experiment 25ph Measurements Buffers And Their Properties
 - The Rise of Digital Reading Experiment 25ph Measurements Buffers And Their Properties
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 25ph Measurements Buffers And Their Properties
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 25ph Measurements Buffers And Their Properties
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 25ph Measurements Buffers And Their Properties
 - Personalized Recommendations
 - Experiment 25ph Measurements Buffers And Their Properties User Reviews and Ratings
 - Experiment 25ph Measurements Buffers And Their Properties and Bestseller Lists
5. Accessing Experiment 25ph Measurements Buffers And Their Properties Free and Paid eBooks

- Experiment 25ph Measurements Buffers And Their Properties Public Domain eBooks
- Experiment 25ph Measurements Buffers And Their Properties eBook Subscription Services
- Experiment 25ph Measurements Buffers And Their Properties Budget-Friendly Options
- 6. Navigating Experiment 25ph Measurements Buffers And Their Properties eBook Formats
 - ePub, PDF, MOBI, and More
 - Experiment 25ph Measurements Buffers And Their Properties Compatibility with Devices
 - Experiment 25ph Measurements Buffers And Their Properties Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 25ph Measurements Buffers And Their Properties
 - Highlighting and Note-Taking Experiment 25ph Measurements Buffers And Their Properties
 - Interactive Elements Experiment 25ph Measurements Buffers And Their Properties
- 8. Staying Engaged with Experiment 25ph Measurements Buffers And Their Properties
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 25ph Measurements Buffers And Their Properties
- 9. Balancing eBooks and Physical Books Experiment 25ph Measurements Buffers And Their Properties
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 25ph Measurements Buffers And Their Properties
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Experiment 25ph Measurements Buffers And Their Properties
 - Setting Reading Goals Experiment 25ph Measurements Buffers And Their Properties
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Experiment 25ph Measurements Buffers And Their Properties
 - Fact-Checking eBook Content of Experiment 25ph Measurements Buffers And Their Properties
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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