

EXPERIMENT: INVESTIGATING PRECIPITATION REACTIONS

PDF

AIM

To compare the solubility of a range of ionic substances through precipitation.

EQUIPMENT

☐ test tubes (6)

Cations (all 0.1 mol L⁻¹) in dropper bottles:

- ☐ sodium nitrate, NaNO₃
- ☐ silver nitrate, AgNO₃
- ☐ lead(II) nitrate, Pb(NO₃)₂
- ☐ ammonium nitrate, NH₄NO₃
- ☐ magnesium sulfate, MgSO₄
- ☐ iron(II) sulfate, FeSO₄
- ☐ copper(II) sulfate, CuSO₄
- ☐ zinc(II) sulfate, ZnSO₄
- ☐ aluminium sulfate, Al₂(SO₄)₃
- ☐ barium chloride, BaCl₂
- ☐ calcium chloride, CaCl₂
- ☐ iron(III) chloride, FeCl₃

☐ test tube rack

Anions (all 0.1 mol L⁻¹) in dropper bottles:

- ☐ potassium nitrate, KNO₃
- ☐ potassium chloride, KCl
- ☐ sodium sulfate, Na₂SO₄
- ☐ sodium carbonate, Na₂CO₃
- ☐ sodium acetate, NaCH₃COO
- ☐ sodium hydroxide, NaOH



SAFETY Wear safety glasses. Lead and barium salts are poisons so avoid contact with skin. If contact occurs wash thoroughly with soap and water. Dispose of as directed by your teacher. Consult "Chemical Safety in Schools" booklet (CSES).

PROCEDURE

- 1 Put a 5 mL of a chemical containing each anion into each of six test tubes. Take care not to mix droppers.
- 2 Add 5 mL of a chemical containing the first cation to each. Record your observations.
- 3 Clean all test tubes thoroughly as incorrect results can arise from contamination of solutions.

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