

Chem 120

$$M_{\text{AgNO}_3} = \frac{m}{V} = \frac{25.523 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} = \frac{1000 \text{ mL}}{1.1} = \boxed{2.3203 \frac{\text{mol}}{\text{L}}}$$

4. Confirm that the molarity is the same molarity that you calculated in step 3.

Yes, the molarity is the same as the molarity calculated in step 3.

Part B: Unknown solutions

Using a method similar to the one used in part A, create and carry out a procedure for determining the molarity of the unknown AgNO_3 solution and the unknown NaCl solution. Record your data and show your calculations.

Unknown AgNO_3

Upon the addition of solid NaCl , solid AgCl was formed.

AgCl formed (g) = 2.3209 g

$$m(\text{AgNO}_3) = \frac{M(\text{AgNO}_3)}{M(\text{AgCl})} \cdot m(\text{AgCl}) = m(\text{AgNO}_3) = \frac{149.87}{143.32} \cdot 2.3209 \text{ g} = \boxed{2.3700 \text{ g}}$$

$$M_{\text{AgNO}_3} = \frac{m}{V} = \frac{2.3700 \text{ g} + 149.87 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} = \frac{1000 \text{ mL}}{1.1} = \boxed{2.5449 \frac{\text{mol}}{\text{L}}}$$

Unknown NaCl

Upon the addition of solid AgNO_3 , solid AgCl was formed.

AgCl formed (g) = 10.7407 g

$$m(\text{NaCl}) = \frac{M(\text{NaCl})}{M(\text{AgCl})} \cdot m(\text{AgCl}) = m(\text{NaCl}) = \frac{58.443}{143.32} \cdot 10.7407 \text{ g} = \boxed{4.4002 \text{ g}}$$

$$M_{\text{NaCl}} = \frac{m}{V} = \frac{4.4002 \text{ g} + 58.443 \frac{\text{g}}{\text{mol}}}{100.00 \text{ mL}} = \frac{1000 \text{ mL}}{1.1} = \boxed{1.7594 \frac{\text{mol}}{\text{L}}}$$

Conclusion questions

1. If you have a company that is using this method to create AgCl , knowing that AgNO_3 is an expensive chemical, what do you think would be a better limiting reagent to use, AgNO_3 or NaCl ? Why?

Exp Limiting Reactant Prelab Answers

Raffaela Di Napoli



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