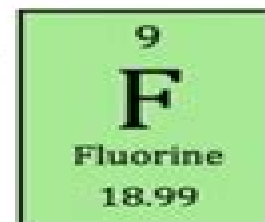
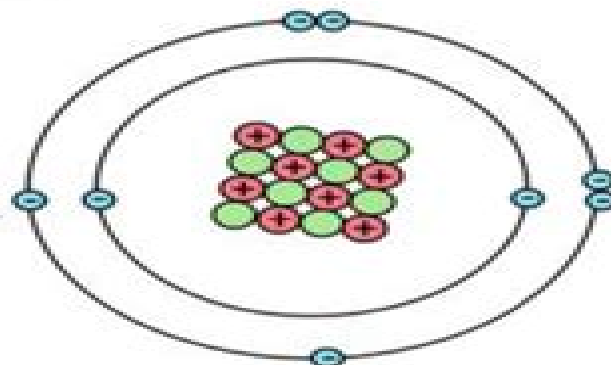


Atomic Theory – Unit Exam

For all of the following multiple-choice questions, choose only one correct answer:

- Which description below accurately describes the current model of the atom?
 - Electrons are distributed around the nucleus and have almost no mass compared to protons
 - Protons, electrons and neutrons are evenly distributed throughout all parts of the atom
 - The nucleus consists of protons, electrons and neutrons
 - The nucleus is made of electrons and protons
- Which of the following is **true** about subatomic particles?
 - Electrons are negatively charged and are the heaviest subatomic particle
 - The mass of a proton nearly equals the mass of a neutron
 - Protons are positively charged and the lightest subatomic particle
 - Neutrons are negatively charged and are the heaviest subatomic particle
- What would be the **mass** of the atom pictured to the right?
 - 7 amu
 - 8 amu
 - 16 amu
 - 24 amu
- How many **valence electrons** are shown in the atom on the right?
 - 2 valence electrons
 - 4 valence electrons
 - 6 valence electrons
 - 8 valence electrons
- Why are atoms electrically **neutral**?
 - The number of neutrons are equal to the number of protons and electrons.
 - The numbers of neutrons and electrons are equal and cancel each other out.
 - The numbers of protons and electrons are equal and cancel each other out.
 - The numbers of protons and neutrons are equal and cancel each other out.
- Based on the symbol to the right, what can be concluded about the element **fluorine**?
 - There are 19 atoms in a molecule of fluorine
 - There are 9 neutrons outside of the fluorine atom's nucleus
 - There are 9 protons in the nucleus of a fluorine atom
 - There are 9 valence electrons in the fluorine atom.



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