

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Electron Configuration Section 5 Answer Key

JG Myers



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Table of Contents Electron Configuration Section 5 Answer Key

1. Understanding the eBook Electron Configuration Section 5 Answer Key
 - The Rise of Digital Reading Electron Configuration Section 5 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Section 5 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Section 5 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Section 5 Answer Key
 - Personalized Recommendations
 - Electron Configuration Section 5 Answer Key User Reviews and Ratings
 - Electron Configuration Section 5 Answer Key and Bestseller Lists
5. Accessing Electron Configuration Section 5 Answer Key Free and Paid eBooks
 - Electron Configuration Section 5 Answer Key Public Domain eBooks
 - Electron Configuration Section 5 Answer Key eBook Subscription Services
 - Electron Configuration Section 5 Answer Key Budget-Friendly Options

6. Navigating Electron Configuration Section 5 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Section 5 Answer Key Compatibility with Devices
 - Electron Configuration Section 5 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Section 5 Answer Key
 - Highlighting and Note-Taking Electron Configuration Section 5 Answer Key
 - Interactive Elements Electron Configuration Section 5 Answer Key
8. Staying Engaged with Electron Configuration Section 5 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Section 5 Answer Key
9. Balancing eBooks and Physical Books Electron Configuration Section 5 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Section 5 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Section 5 Answer Key
 - Setting Reading Goals Electron Configuration Section 5 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Section 5 Answer Key
 - Fact-Checking eBook Content of Electron Configuration Section 5 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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