

## Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- 1) nitrogen       $1s^2 2s^2 2p^3$
- 2) copper       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
- 3) chlorine       $1s^2 2s^2 2p^6 3s^2 3p^5$
- 4) calcium       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- 5) lead       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

In the space below, write the noble gas notation electron configurations of the following elements:

- 6) barium       $[Xe] 6s^2$
- 7) gold       $[Xe] 4f^{14} 5d^{10} 6s^1$
- 8) platinum       $[Xe] 4f^{14} 5d^9 6s^1$
- 9) xenon       $[Kr] 4d^{10} 5s^2 p^6$
- 10) calcium       $[Ar] 4s^2$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^3$       phosphorous
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$       ruthenium
- 13)  $[Kr] 4d^{10} 5s^2 p^5$       iodine
- 14)  $[Xe] 4f^{14} 5d^{10} 6s^2$       mercury
- 15)  $[Rn] 7s^2$       radium

Determine which of the following electron configurations are not correct:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$        $3p^6$  should be a 6
- 17)  $1s^2 2s^2 2p^6 3s^2 2p^5$       1s only holds 2 and there aren't two "2p" orbitals
- 18)  $[Cs] 5f^{14} 6d^1 7s^2$       Cs is not a noble gas
- 19)  $[Kr] 4d^{10} 5s^2 p^4$       Correct
- 20)  $[Ne]$  for Neon      Can't use a noble gas to shortcut itself

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**LM Reder**



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