

CHM 101

Electron Structure Worksheet

Objectives:

- Explain the nature of light and how it reveals the electron structure of the atom.
- Write the electron structure of elements
- Relate electron structure to the periodic table

1. Light is a form of what type of energy?
 - Give 2 examples of other forms of this type of energy.
2. What is the meaning of wavelength and frequency of light?
3. White light produces a continuous spectrum. Elements that are excited produce a line spectrum. What is a line spectrum?
4. Different metals emit different colored light in flame tests. This light originates from inside the metal atoms. Explain what was happening inside these atoms that causes the various colors of light. Be specific
5. What is an orbital?
6. List the orbitals that exist in the following energy levels. Indicate the number of these orbitals.
 $n = 1$, $n = 2$, $n = 3$, $n = 4$
7. Indicate the electron configuration of these elements using the shorthand notation and the orbital diagrams.



For example - helium: $1s^2$, $1s$

Elements: Li, C, Mg, Fe, Kr

8. In this periodic chart, indicate the outer orbitals being filled with electrons in the rectangles.

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9. Why do elements in the same group have similar properties?

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L Manion

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