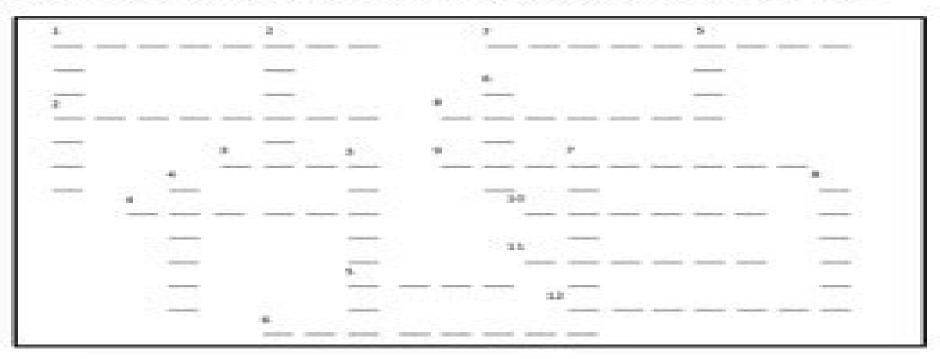
# Electron Configuration Crossword

Directions: Each entry in the puzzle has an electron configuration as a clue. Write the name of each element with one letter on each line.



#### Across

- 1. 1s22s22p43s23p44s23d5
- 2. 192292205392302
- 3. [Hell2s/22p\*
- 4. 1s22s22p33s23p44s23d1
- 5.00e16s34835d10
- 6. 1s22s22p43s23p44s23d1
- 7. 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>
- B. 1s22s22p3s23p4s23d24p4
- 9. 197297201397301497
- 10. 15\*25\*20\*35\*30\*
- 11.00/15s34d38
- 12. [8n]7s3545d3

#### Displaying

- INel3s\*3p\*4s\*
- [Kr]5s<sup>1</sup>4d<sup>20</sup>5p<sup>1</sup>
- 3. 18<sup>2</sup>28<sup>2</sup>29<sup>3</sup>
- 4. [Art]45/3d\*\*
- 1s<sup>2</sup>2s<sup>2</sup>2p<sup>3</sup>3s<sup>2</sup>3p<sup>4</sup>4s<sup>2</sup>3d<sup>8</sup>
- 1s<sup>1</sup>2s<sup>1</sup>2p<sup>6</sup>3s<sup>1</sup>3p<sup>6</sup>
- 7. 1s<sup>2</sup>2s<sup>1</sup>2o<sup>4</sup>3s<sup>2</sup>3o<sup>4</sup>
- [D0e] 65246745d6

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Electron shell List of elements with electrons per shell Periodic table Layout Rutherfordium Chemical Antimony Group 8 element Lanthanum Chemical properties Extended periodic table Eka superactinides Manganese Promethium Physical properties and much more The Periodic Table Based on Electron Configuration Raymond Rivet, 1979 Elements and Their Electron Configurations Kroutil Artists, 2004-01-01 With an introduction to the elements in character students can build all 109 atoms for the named chemical elements This multisensory approach to teaching is producing a 90% class average for the students who use it The first step in understanding the atomic structure of the elements starts here Book includes a periodic table and 109 color electron configuration stickers **Electronic Configuration: A Formula** Handbook N.B. Singh, Electronic Configuration A Formula Handbook is a concise and indispensable guide for understanding the arrangement of electrons in atoms and molecules This handbook provides clear and easy to follow formulas and rules for determining electronic configurations enabling readers to guickly and accurately predict the distribution of electrons in various atomic and molecular systems Whether you re a student studying chemistry or a professional in the field this book serves as a valuable reference for mastering electronic configurations and their implications in chemical bonding and reactivity **Electron Configuration as the Basis of the Periodic Table William** Atomic and Molecular Electronic Configuration Revisited Harold J. Teague, 2014-06-11 In AMEC chapter Fav Luder.1943 one you will find a structure for O2 supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O2 fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O2 but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the NO and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O2 OF F2 and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE

method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an elect added	ron is

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