

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the hypochlorite ion is reduced? (Note: the reaction is not balanced.)



144: 1 (C) 145: 2 (B) 146: 3 (B) 147: 4 (B)



PG-2: What is the balanced half reaction of the species in reaction 1, when nitrogen is reduced in acidic solution?



PG-3: What is the coefficient for water when the reaction is balanced in acidic solution?

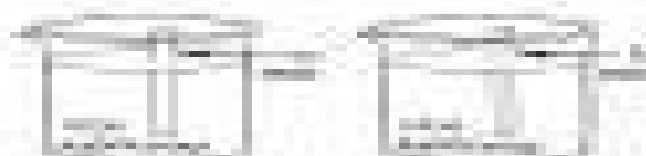


unbalanced



PG-4: Which reaction occurs in the cell when the phones are connected from their recharged state?

The standard reduction potential of zinc is $+1.26 \text{ V}$.
The standard reduction potential of iron is $+0.77 \text{ V}$.



The cell is to do work, releasing energy. The reduction half reaction is $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$. The oxidation half reaction is $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$. The overall reaction is $\text{Zn} + \text{Fe}^{2+} \rightarrow \text{Zn}^{2+} + \text{Fe}$.

For a reaction to be spontaneous, the cell potential must be positive. The cell potential is $E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}} = 0.77 \text{ V} - 1.26 \text{ V} = -0.49 \text{ V}$. The negative value indicates the reaction is non-spontaneous. The reaction is spontaneous when the cell potential is positive.

Electrochemistry Answers Study Guide

Rajeevan N



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