## Electrochemical Cells | 17 Pre-lab and Thermodynamics

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

- Define the following terms: faraday, salt bridge, anode, cathode, voltaic cell, and electrolytic cell.
- Write a chemical equation for the reaction that occurs in the following cell: Cu | Cu2+ (aq) | Ag+ (aq) | Ag-
- Given the following E's, calculate the standard-cell potential for the cell in question 2.

$$Cu^{2+}(aq) + 2e^{-} \longrightarrow Cu(s)$$
  $E^{+} = +0.34 \text{ V}$   
 $Ag^{+}(aq) + e^{-} \longrightarrow Ag(s)$   $E^{-} = +0.80 \text{ V}$ 

Calculate the voltage of the following cell:

Calculate the cell potential, the equilibrium constant, and the free-energy change for

$$Ca(s) + Mn^{2+}(oq)(1 M) \Longrightarrow Ca^{2+}(oq)(1 M) + Mn(s)$$

given the following Eo values:

$$Ca^{2+}(aq)+2e^{-}\longrightarrow Ca(s)$$
  $E^{0}=-2.87 \text{ V}$   
 $Mn^{2+}(aq)+2e^{-}\longrightarrow Mn(s)$   $E^{0}=-1.18 \text{ V}$ 

# **Electrochemical Cell Pre Lab Answers**

**Richard Bailey** 



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