

63. The metal that cannot be obtained by electrolysis of an aqueous solution of its salts is :

- (1) Cu (2) Cr (3) Ag (4) Ca

63. (4)
Factual

64. Resistance of 0.2 M solution of an electrolyte is 50 Ω. The specific conductance of the solution is 1.4 S m^{-1} . The resistance of 0.5 M solution of the same electrolyte is 200 Ω. The molar conductivity of 0.5 M solution of the electrolyte in $\text{S m}^2 \text{ mol}^{-1}$ is :

- (1) 5×10^2 (2) 5×10^3 (3) 5×10^{-4} (4) 5×10^{-5}

64. (2)

$$\Lambda_m = \frac{k (\text{s/m})}{1000 \times \text{Molarity} (\text{mol}^{-1} \text{L})} = \frac{1}{4 \times 1000 \times 0.5}$$

$$= 5 \times 10^{-4} \text{ Sm}^2 \text{ mol}^{-1}$$

65. CaCl crystallises in body centred cubic lattice. If 'a' is its edge length then which of the following expressions is correct ?

- (1) $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{\sqrt{3}}{2} a$ (2) $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \sqrt{3} a$ (3) $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = 3a$ (4) $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{3a}{2}$

65. (1)
Factual

66. Consider separate solutions of 0.500 M $\text{C}_2\text{H}_5\text{OH}(\text{aq})$, 0.100 M $\text{Mg}_3(\text{PO}_4)_2(\text{aq})$, 0.250 M $\text{KBr}(\text{aq})$ and 0.125 M $\text{Na}_3\text{PO}_4(\text{aq})$ at 25°C. Which statement(s) is/are true about these solutions, assuming all salts to be strong electrolytes ?

- (1) 0.125 M $\text{Na}_3\text{PO}_4(\text{aq})$ has the highest osmotic pressure.
 (2) 0.500 M $\text{C}_2\text{H}_5\text{OH}(\text{aq})$ has the highest osmotic pressure.
 (3) They all have the same osmotic pressure.
 (4) 0.100 M $\text{Mg}_3(\text{PO}_4)_2(\text{aq})$ has the highest osmotic pressure.

66. (3)
All solutions have same osmotic pressure.

67. In which of the following reactions H_2O_2 acts as a reducing agent ?

- (a) $\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O}$ (b) $\text{H}_2\text{O}_2 + 2\text{e}^{-} \rightarrow \text{O}_2 + 2\text{H}^{+}$
 (c) $\text{H}_2\text{O}_2 + 2\text{e}^{-} \rightarrow 2\text{OH}^{-}$ (d) $\text{H}_2\text{O}_2 + 2\text{OH}^{-} + 2\text{e}^{-} \rightarrow \text{O}_2 + 2\text{H}_2\text{O}$
 (1) (a), (c) (2) (b), (d) (3) (a), (b) (4) (c), (d)

67. (2)

$$\text{H}_2\text{O}_2 \rightarrow \text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-}$$

$$\text{H}_2\text{O}_2 + 2\text{OH}^{-} \rightarrow \text{O}_2 + 2\text{H}_2\text{O} + 2\text{e}^{-}$$

68. In $\text{S}_{\text{N}}2$ reactions, the correct order of reactivity for the following compounds :

- CH_3Cl , $\text{CH}_3\text{CH}_2\text{Cl}$, $(\text{CH}_3)_2\text{CHCl}$ and $(\text{CH}_3)_3\text{CCl}$ is :
- (1) $\text{CH}_3\text{CH}_2\text{Cl} > \text{CH}_3\text{Cl} > (\text{CH}_3)_2\text{CHCl} > (\text{CH}_3)_3\text{CCl}$
 (2) $(\text{CH}_3)_2\text{CHCl} > \text{CH}_3\text{CH}_2\text{Cl} > \text{CH}_3\text{Cl} > (\text{CH}_3)_3\text{CCl}$
 (3) $\text{CH}_3\text{Cl} > (\text{CH}_3)_2\text{CHCl} > \text{CH}_3\text{CH}_2\text{Cl} > (\text{CH}_3)_3\text{CCl}$
 (4) $\text{CH}_3\text{Cl} > \text{CH}_3\text{CH}_2\text{Cl} > (\text{CH}_3)_2\text{CHCl} > (\text{CH}_3)_3\text{CCl}$

68. (4)
Order of reactivity of alkyl halide towards $\text{S}_{\text{N}}2$ reaction is $\text{CH}_3\text{X} > 1^\circ > 2^\circ > 3^\circ$

69. The octahedral complex of a metal ion M^{3+} with four monodentate ligands L_1 , L_2 , L_3 and L_4 absorb wavelengths in the region of red, green, yellow and blue, respectively. The increasing order of ligand strength of the four ligands is :

- (1) $\text{L}_1 < \text{L}_2 < \text{L}_4 < \text{L}_3$ (2) $\text{L}_1 < \text{L}_2 < \text{L}_3 < \text{L}_4$ (3) $\text{L}_4 < \text{L}_2 < \text{L}_3 < \text{L}_1$ (4) $\text{L}_1 < \text{L}_3 < \text{L}_2 < \text{L}_4$

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