

NOTES - Unit 4: Bonding

PART 1: Ionic and Covalent Bonding

CHEMICAL BONDS

- A chemical bond is an attraction between 2 atoms or ions.
- Bonding occurs because it _____ of the system.
- Three broad classifications (general rule of thumb)
 - Ionic (_____ + _____)
 - Covalent (_____ + _____)
 - Metallic (_____ + _____)
- _____: the power of an atom in a molecule to attract electrons to itself.
 - The greater an atom's electronegativity, the greater its ability to attract _____.
 - In a compound, the element with the greater electronegativity will be the more _____ species.
 - The element with _____ electronegativity is sometimes referred to as the more " _____ " element. This will become the more positive species in the bond.
- _____ bonds exist between atoms with low electronegativities & those with high electronegativities. In general, between metals and nonmetals.
- _____ bonds exist between dissimilar atoms with different electronegativities.
- _____ bonds exist between identical nonmetallic atoms or nonmetallic atoms with similar electronegativities.
- Atoms react with each other in chemical reactions in a quest to have complete outer electron energy levels like the _____.
 - Octet Rule: Atoms tend to gain, lose, or share electrons until they are surrounded by _____ valence electrons.
 - Duet Rule: _____ and _____ want to have 2 electrons. Recall that the first energy level only holds two electrons.

IONIC BONDING

- Different texts report different "cutoff" values, but according to IB, _____ for ionic bonds.
- The greater the difference in electronegativity between the atoms, the more _____ the bond.
- Ionic bonds form between metals on the _____ of the periodic table and elements in groups _____ (or polyatomic ions attracted to other ions)
 - Note: group 4 elements (C and Si) do not form ionic bonds. They tend to form giant molecular structures (network solids) or simple molecules.
- Example: NaCl
 - Sodium (Na) has 11 electrons (but only one valence electron): _____
 - By losing this electron Na becomes Na^+ : _____ or _____ which has a full outer shell.
 - Chlorine (Cl) has 17 electrons (seven valence electrons): _____

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A Loxley



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