#### Question 33 (4 marks)

A student adds 1.17 g of Al(OH)<sub>3</sub>(s) to 0.500 L of 0.100 mol L<sup>-1</sup> HCl(ag).

Calculate the pH of the resulting solution. Assume that the volume of the resulting solution is 0.500 L.

1) write a balanced chemical equation

A1 (0H)3 (5) + 3 HC (109) - A1 C13 (49 1+ 3 H2 OC)

(2) Find no. of mole > of QH and H+  $n(AI(OH)_3) = \frac{m}{mm} = \frac{1}{24.98 + 3(16 + 1.008)} = 0.01499.$ 

 $n(0H^{-}) = 3 \times n(AI(0H)_{3}) = 0.04499...mp1$ 

 $N(H(I) = C \times V = 0.5 \times 0.1 = 0.05 \text{ mol}$ 

 $n(H^{+}) = n(HCI) = 0.05 mel$ 

3) Determine which reagent is in excess and now much

D. 05 - 0.04499 .. = 0.00500 ... MD)

1 at 1 the [H+] and thus find pH.

PH = -109 10 CH+7 = 2.000 (3 dp)

Please turn over

# **Hsc Chemistry 6 User Guide**

Jim Stamell

# **Hsc Chemistry 6 User Guide:**

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