

### Vocab:

64. Unsaturated

65. Saturated  
maximum amt. of solute dissolved at that temp.

66. Solubility  
the amount (g) of solute that can dissolve at a certain temp.

67. Rate of dissolution  
how fast a solute can dissolve

68. Molarity  
concentration  $M = \frac{\text{mol solute}}{1 \text{ Liter}}$ 

69. Supersaturated  
more than max amt. of solute dissolved in solvent

70. Electrolyte  
- must be soluble  
- must be ionic  
can conduct electricity

71. List 3 ways to increase the rate of dissolution of a solid:

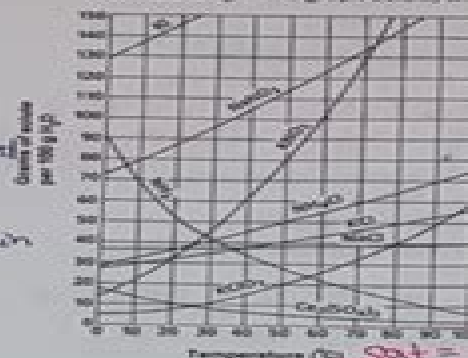
1. increase surface area (crush)
2. increase temperature
3. stir

### SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

Soluble compounds contain	Common exceptions
$\text{Cl}^-$ , $\text{Br}^-$ , $\text{I}^-$ , $\text{NO}_3^-$	None
$\text{SO}_4^{2-}$	None
$\text{CO}_3^{2-}$	None
$\text{OH}^-$	None
$\text{S}^{2-}$	None
$\text{PO}_4^{3-}$	None
$\text{C}_2\text{O}_4^{2-}$	None
$\text{F}^-$	Compounds of $\text{Ag}^+$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$
$\text{SCN}^-$	Compounds of $\text{Ag}^+$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$
$\text{CN}^-$	Compounds of $\text{Ag}^+$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$
$\text{C}_2\text{O}_4^{2-}$	Compounds of $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$

Insoluble compounds contain	Common exceptions
$\text{CO}_3^{2-}$	Compounds of $\text{Na}^+$ and the alkali metal cations
$\text{PO}_4^{3-}$	Compounds of $\text{Na}^+$ and the alkali metal cations
$\text{C}_2\text{O}_4^{2-}$	Compounds of $\text{Na}^+$ and the alkali metal cations
$\text{OH}^-$	Compounds of $\text{Na}^+$ and the alkali metal cations, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , and $\text{Ba}^{2+}$
$\text{S}^{2-}$	Compounds of $\text{Na}^+$ , the alkali metal cations, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , and $\text{Ba}^{2+}$

72. According to the graph below, which substance maintains almost constant solubility over the given temperature range?



NaCl

73. Which substance decreases in solubility with increasing temperature?

$\text{NH}_3$ ,  $\text{Ce}_2(\text{SO}_4)_3$

74. What is the solubility of  $\text{KNO}_3$  at  $60^\circ\text{C}$ ?

$\sim 102\text{g}$

75. A solution containing 50 g  $\text{CaCl}_2$  in 100 g  $\text{H}_2\text{O}$  is prepared at  $10^\circ\text{C}$ . Is the solution saturated, unsaturated or supersaturated?

(Not on graph :))

76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
$\text{NaNO}_3$	✓		$\text{NaNO}_3 \rightarrow \text{Na}^+ + \text{NO}_3^-$	✓	2
$\text{Na}_2\text{CO}_3$	✓		$\text{Na}_2\text{CO}_3 \rightarrow 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
$\text{NaCl}$	✓		$\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \rightarrow \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
$\text{CaCO}_3$		✓	does not dissociate	NO	0

Molarity (M) is the number of moles of solute in one liter of solution

$$\text{molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

Dilution: the action of making a stock solution less concentrated

$$M_1V_1 = M_2V_2$$

77. Calculate the Molarity of a solution that contains 100 grams of  $\text{NaCl}$  in 5.00L of solution [the molar mass of  $\text{NaCl}$  is 58.44 g/mol]

Convert g  $\rightarrow$  mol

$$\frac{100 \text{ g NaCl}}{58.44 \text{ g/mol}} = 1.71 \text{ mol}$$

$$M = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution

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