

SIDEWALL POWER VENTER KIT

Model: SWG, SWGIII & SWG Stainless Series

*Patented



TYPICAL VENTING SYSTEM COMPONENTS

- 1 - SWG Series Power Venter
- 1 - CK Series Control Kit (sold separately)

OPTIONAL SYSTEM COMPONENTS

SWG SERIES THROUGH-WALL EXTENSION KIT

For installation in wall thickness over 8 inches. Models PEX-4 through PEX-8 are available.

FOR MOST MULTIPLE HEATING EQUIPMENT SYSTEMS

One CK Series Control Kit for each appliance. Except for a 24V gas furnace or boiler and a 30mW water heater multiple venting system, use the CK-60 Series Control Kit.

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GENERAL SYSTEM INFORMATION

Designed for operation with natural gas, LP gas and #2 fuel oil appliances.

1. The thermostat (wall thermostat or aquastat) calls for heat and energizes a relay which activates the power venter. After the venter motor has come up to speed, the pressure switch closes. This closes the circuit to the burner and allows the burner to fire.
2. For millivolt controlled water heaters using the CK-20 Series Control Kit, the gas valve pressure switch activates the power venter at the same time as the burner fires.
3. After the heating equipment has been satisfied, the thermostat circuit will open and de-activate the burner and power venter circuit.
4. For venting systems equipped with a post purge device, the power venter operates for a period of time after the burner has shut off to purge remaining fuel gases.

DO NOT DESTROY

THESE INSTRUCTIONS MUST REMAIN WITH EQUIPMENT



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Yicheng Fang



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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = 57\text{ kJ mol}^{-1}$ eg 2 $\text{HNO}_3\text{(aq)} + \text{KOH(aq)} \rightarrow \text{KNO}_3\text{(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = 57\text{ kJ mol}^{-1}$

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