

10. Factors Affecting the Rate of a Chemical Reaction

What you will accomplish in this experiment

You'll investigate *two* of the *four* factors that affect the rate of a chemical reaction.

Specifically, you'll study the impact of:

- Increasing the *concentration* of each reactant (increasing the number of *moles of reactant per liter of solution*).
- Increasing the *temperature* of the reaction by heating.

You'll determine the rate of a reaction by measuring the decrease in reactant concentration over a period of time (as it is being converted into product). You'll then need to make a conclusion as to whether reaction rate *increases* or *decreases* when the *concentration* of each reactant is increased, and when the overall temperature of the reaction is increased.

Concepts you need to know to be prepared

The study of the *rate* of a chemical reaction is called *chemical kinetics*.

The rate of a chemical reaction can be studied by measuring either:

- How quickly a *reactant* is used up, or
- How quickly a *product* is formed.

More specifically, the *change in concentration* of a reactant (or of a product) is measured *over a certain period of time* as the reaction is occurring.

For example, the figure to the left below illustrates how one might measure the rate of decomposition of hydrogen peroxide, H_2O_2 . The hydrogen peroxide molecules decompose to produce water and oxygen gas.



The change in mass as the bubbles of oxygen gas diffuse out of the reaction flask provides a way to measure how quickly the product is being formed. The graph to the right shows the increase in oxygen concentration (as well as the decrease in hydroxide peroxide concentration) over a period of time.



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