

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (1) (10-11 P. 41) Explain a new scientific theory of atoms (the modern atomic theory) by describing changes in the atomic model over time, including those changes based on experimental observations. (10-11 P. 42) Explain the scientific theory of atomic orbitals (model for atomic theory) by describing the scientific changes in terms of quantum mechanical calculations, including wave-particle duality and how it helps explain all their mass, electric charges, and locations within the atom. (10-11 P. 42) Explain properties of atoms and their patterns in the periodic table in terms of arrangements of their electrons, related to Pauli exclusion principle, Hund's rule, and the four fundamental forces (gravitational, electromagnetic, weak force, strong force).

ANSWER KEY

Directions: Match the atomic theory selected on the correct card(s).

Model(s)	Model(s)	Atomic Theory's Basic Principles
(C)	Democritus (460 BC)	A. Everything is made of atoms that cannot be created, divided, or destroyed
(A)	Bohr (1913)	B. When Electrons Absorb energy, usually in packets, charged matter is accelerated and travels throughout the positive material
(B)	Thomson (1897)	C. "atomic indivisible (uncuttable)"
(F)	Rutherford (1911)	D. Group Number (Valence Electrons)
(D)	Moseley (1913)	E. Electrons Cloud
(E)	Schrodinger (1926)	F. Atoms randomly occupy space the subatomic particles must

Directions: Use the periodic table to answer the following questions:

- (10-11)** **(10-11)** The number of protons equals the number of electrons.
- atomic mass minus the atomic number will equal the **Number of neutrons**.
- (10-11)** **(10-11)** **(10-11)** protons are positive (positively charged) electrons are negative (negatively charged).

Electrons In Atoms Answer Key Guided

M Walker



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