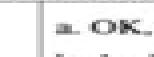


8 • Electron Configurations & Periodicity**ANSWERS TO STUDY QUESTIONS**

1. Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
- phosphorus [Ne] 3s² 3p³
 - nickel [Ar] 3d⁸ 4s²
 - osmium [Xe] 4f¹⁴ 5d⁶ 6s²
 - californium [Rn] 5f¹⁰ 7s²
 - titanium [Ar] 3d² 4s²
2. Which orbital is filled following these orbitals?
- 3d is followed by 4p
 - 4s is followed by 3d
 - 5p is followed by 6s
 - 5f is followed by 6d
3. How many electrons can be accommodated in
- a d subshell 10; two in each of 5 orbitals
 - a set of f orbitals 14; two in each of 7 orbitals
 - the n = 4 shell 32; 2z²
 - the 7s orbital 2
 - a p_x orbital? 2
4. What is wrong with the following ground state electron configurations?
- | | |
|---|--|
| a.


 | a. OK. 4s is filled before 3d. |
| b.


 | b. 4s should be filled before 4p subshell. |
| c.


 | c. 4s cannot have 2 electrons with same spin. |
| d.


 | d. OK, stability of half-filled 3d subshell. |
| e.


 | e. 4p should first fill one orbital at a time. |
5. How many unpaired electrons are there in
- a nitrogen atom 3 unpaired electrons (2s² 2p³)
 - an iodine atom 1 unpaired electrons (5s² 5p⁵)
 - a nickel (II) cation 2 unpaired electrons (3d⁸)
 - an oxide ion no unpaired electrons (closed shell [Ne]⁸)

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Judd E. Hollander

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