

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
also might be written (due to the close energy of d and f and similarities of elements versus predictable)
memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
also might be written (due to the close energy of d and f and similarities of elements versus predictable)
memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium note: no d's, so the p. table referenced has Ac as 5f
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Electron Configuration Ws 1 Answer Key

Michael Brown



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Table of Contents Electron Configuration Ws 1 Answer Key

1. Understanding the eBook Electron Configuration Ws 1 Answer Key
 - The Rise of Digital Reading Electron Configuration Ws 1 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Ws 1 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Ws 1 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Ws 1 Answer Key
 - Personalized Recommendations
 - Electron Configuration Ws 1 Answer Key User Reviews and Ratings

- Electron Configuration Ws 1 Answer Key and Bestseller Lists
- 5. Accessing Electron Configuration Ws 1 Answer Key Free and Paid eBooks
 - Electron Configuration Ws 1 Answer Key Public Domain eBooks
 - Electron Configuration Ws 1 Answer Key eBook Subscription Services
 - Electron Configuration Ws 1 Answer Key Budget-Friendly Options
- 6. Navigating Electron Configuration Ws 1 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Ws 1 Answer Key Compatibility with Devices
 - Electron Configuration Ws 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Ws 1 Answer Key
 - Highlighting and Note-Taking Electron Configuration Ws 1 Answer Key
 - Interactive Elements Electron Configuration Ws 1 Answer Key
- 8. Staying Engaged with Electron Configuration Ws 1 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Ws 1 Answer Key
- 9. Balancing eBooks and Physical Books Electron Configuration Ws 1 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Ws 1 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Ws 1 Answer Key
 - Setting Reading Goals Electron Configuration Ws 1 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Ws 1 Answer Key
 - Fact-Checking eBook Content of Electron Configuration Ws 1 Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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