

Electron Configuration Practice
Chemistry

Name: _____

Due Date: _____

How to write an electron configuration:

- Determine the total number of electrons to be represented.
- Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- The sum of the superscripts should equal the total number of electrons. Example:
 $12\text{Mg } 1s^2 2s^2 2p^6 3s^2$

I. Configuration Writing Practice

- Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2 2s^2 2p^6 3s^1$

2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$

4. U: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3 6d^1$

5. N: $1s^2 2s^2 2p^3$

* 6. Ag: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$

7. Ti: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

8. Ce: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^1 5d^1$

9. Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$

10. Hg: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Electron Configuration Practice With Answers

L Towner



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Table of Contents Electron Configuration Practice With Answers

1. Understanding the eBook Electron Configuration Practice With Answers
 - The Rise of Digital Reading Electron Configuration Practice With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Practice With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Practice With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Practice With Answers
 - Personalized Recommendations
 - Electron Configuration Practice With Answers User Reviews and Ratings
 - Electron Configuration Practice With Answers and Bestseller Lists
5. Accessing Electron Configuration Practice With Answers Free and Paid eBooks
 - Electron Configuration Practice With Answers Public Domain eBooks
 - Electron Configuration Practice With Answers eBook Subscription Services
 - Electron Configuration Practice With Answers Budget-Friendly Options
6. Navigating Electron Configuration Practice With Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Electron Configuration Practice With Answers Compatibility with Devices
 - Electron Configuration Practice With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Practice With Answers
 - Highlighting and Note-Taking Electron Configuration Practice With Answers
 - Interactive Elements Electron Configuration Practice With Answers
 8. Staying Engaged with Electron Configuration Practice With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Practice With Answers
 9. Balancing eBooks and Physical Books Electron Configuration Practice With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Practice With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Electron Configuration Practice With Answers
 - Setting Reading Goals Electron Configuration Practice With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Electron Configuration Practice With Answers
 - Fact-Checking eBook Content of Electron Configuration Practice With Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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