

8 • Electron Configurations & Periodicity

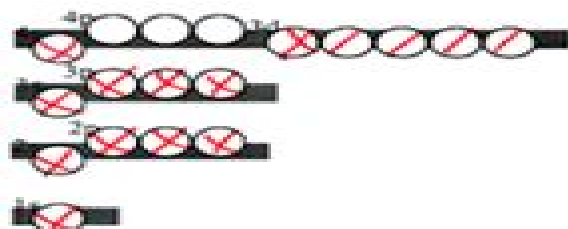
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^4$
Ar $[Ne] 3s^2 3p^4$
- Mg $1s^2 2s^2 2p^6 3s^2$
Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
Li $[He] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Electron Configuration Practice Answer Key Table

Lingjun Ying



Electron Configuration Practice Answer Key Table:

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Table of Contents Electron Configuration Practice Answer Key Table

1. Understanding the eBook Electron Configuration Practice Answer Key Table
 - The Rise of Digital Reading Electron Configuration Practice Answer Key Table
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Practice Answer Key Table
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Practice Answer Key Table
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Practice Answer Key Table
 - Personalized Recommendations
 - Electron Configuration Practice Answer Key Table User Reviews and Ratings
 - Electron Configuration Practice Answer Key Table and Bestseller Lists
5. Accessing Electron Configuration Practice Answer Key Table Free and Paid eBooks

- Electron Configuration Practice Answer Key Table Public Domain eBooks
- Electron Configuration Practice Answer Key Table eBook Subscription Services
- Electron Configuration Practice Answer Key Table Budget-Friendly Options
- 6. Navigating Electron Configuration Practice Answer Key Table eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Practice Answer Key Table Compatibility with Devices
 - Electron Configuration Practice Answer Key Table Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Practice Answer Key Table
 - Highlighting and Note-Taking Electron Configuration Practice Answer Key Table
 - Interactive Elements Electron Configuration Practice Answer Key Table
- 8. Staying Engaged with Electron Configuration Practice Answer Key Table
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Practice Answer Key Table
- 9. Balancing eBooks and Physical Books Electron Configuration Practice Answer Key Table
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Practice Answer Key Table
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Practice Answer Key Table
 - Setting Reading Goals Electron Configuration Practice Answer Key Table
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Practice Answer Key Table
 - Fact-Checking eBook Content of Electron Configuration Practice Answer Key Table
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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