

Model 2 – Ground State Orbital Diagrams and Electron Configurations



Hydrogen
 $1s^1$



Carbon
 $1s^2 2s^2 2p^2$



Oxygen
 $1s^2 2s^2 2p^4$



Sodium
 $1s^2 2s^2 2p^6 3s^1$



Aluminium
 $1s^2 2s^2 2p^6 3s^2 3p^1$



Phosphorus
 $1s^2 2s^2 2p^6 3s^2 3p^3$

6. Examine the orbital diagrams and electron configurations in Model 2. Match each symbol below with its meaning.

___ a. 


I. Single electron

___ b. 

II. Pair of electrons with opposite spins

___ c. 

III. Atomic orbital (region of space where an electron is likely to be found)

___ d. 

IV. Sublevel (set of orbitals having equivalent energy)

___ e. $1s^2 2s^2 2p^4$

V. Electron configuration

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M Mark



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