

# 8 • Electron Configurations & Periodicity

## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e <sup>-</sup> s: <u>18</u>	Element: Mg # of e <sup>-</sup> s: <u>12</u>	Element: N # of e <sup>-</sup> s: <u>7</u>	Element: Li # of e <sup>-</sup> s: <u>3</u>	Element: P # of e <sup>-</sup> s: <u>15</u>	Element: Cl # of e <sup>-</sup> s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^6$   
 Ar  $[Ne] 3s^2 3p^6$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
 Mg  $[Ne] 3s^2$
- N  $1s^2 2s^2 2p^3$   
 N  $[He] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
 Li  $[He] 2s^1$
- P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
 P  $[Ne] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
 Cl  $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$

Fe  $[Ar] 3d^6 4s^2$

# Electron Arrangements In Atoms Worksheet Answers

**Chandan Sukumar Sengupta**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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