

C_2H_4 $8+4$ # of electrons = <u>12</u> $\begin{array}{c} \text{H} & & \text{H} \\ & & \\ \text{H} : \text{C} : & : & \text{C} : \text{H} \end{array}$	C_2F_4 $8+28$ # of electrons = <u>36</u> $\begin{array}{c} : \ddot{\text{F}} : & : \ddot{\text{F}} : \\ & & \\ \text{C} : & : & \text{C} \\ & & \\ : \ddot{\text{F}} : & : \ddot{\text{F}} : \end{array}$	CO $4+6$ # of electrons = <u>10</u> $: \text{C} : : : \text{O} :$	O_2 $6+6$ # of electrons = <u>12</u> $\ddot{\text{O}} : : \ddot{\text{O}}$
CO_2 $4+6+6$ # of electrons = <u>16</u> $\ddot{\text{O}} : : \text{C} : : \ddot{\text{O}}$	C_2H_2 (H C C H) $8+2$ # of electrons = <u>10</u> $\text{H} : \text{C} : : \text{C} : \text{H}$	N_2 $5+5$ # of electrons = <u>10</u> $: \text{N} : : : \text{N} :$	HCN $1+4+5$ # of electrons = <u>10</u> $\text{H} : \text{C} : : : \text{N} :$
CN^- $4+5+1$ # of electrons = <u>10</u> $[: \text{C} : : : \text{N} :]^-$	SO_4^{2-} $6+24+2$ # of electrons = <u>32</u> $\left[\begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{S} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \end{array} \right]^{2-}$	PO_4^{3-} $5+24+3$ # of electrons = <u>32</u> $\left[\begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{P} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \end{array} \right]^{3-}$	ClO_3^- $7+18+1$ # of electrons = <u>26</u> $\left[\begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{Cl} : \ddot{\text{O}} : \end{array} \right]^-$
CO_3^{2-} $4+18+2$ # of electrons = <u>24</u> $\left[\begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{C} : \ddot{\text{O}} : \end{array} \right]^{2-}$ <p><i>Handwritten notes:</i> $\text{O}=\text{C}=\text{O}$ with arrows pointing to the double bonds.</p>	NO_3^- $5+18+1$ # of electrons = <u>24</u> $\left[\begin{array}{c} : \ddot{\text{O}} : \\ : \ddot{\text{O}} : \text{N} : \ddot{\text{O}} : \end{array} \right]^-$ <p><i>Handwritten notes:</i> $\text{O}=\text{N}=\text{O}$ with arrows pointing to the double bonds.</p>	SO_2 $6+12$ # of electrons = <u>18</u> $\begin{array}{c} : \ddot{\text{O}} : \text{S} : \ddot{\text{O}} : \\ \updownarrow \\ \text{O} : \text{S} : \text{O} : \end{array}$	O_3 (O O O) $6+6+6$ # of electrons = <u>18</u> $\begin{array}{c} \ddot{\text{O}} : : \ddot{\text{O}} : \ddot{\text{O}} : \\ \updownarrow \\ \ddot{\text{O}} : \ddot{\text{O}} : : \ddot{\text{O}} : \end{array}$
SF_6 $6+42$ # of electrons = <u>48</u> $\begin{array}{c} & \text{F} & & \text{F} & & \\ & & & & & \\ \text{F} & : & \text{S} & : & \text{F} \\ & & & & & \\ & \text{F} & & \text{F} & & \end{array}$	XeF_4 $8+28$ # of electrons = <u>36</u> $\begin{array}{c} & \text{F} & & \\ & & & \\ \text{F} & : & \text{Xe} & : & \text{F} \\ & & & \\ & \text{F} & & \end{array}$	PCl_5 $5+35$ # of electrons = <u>40</u> $\begin{array}{c} & \text{Cl} & & \\ & & & \\ \text{Cl} & : & \text{P} & : & \text{Cl} \\ & & & \\ & \text{Cl} & & \end{array}$	SeF_4 $6+28$ # of electrons = <u>34</u> $\begin{array}{c} & \text{F} & & \\ & & & \\ \text{F} & : & \text{Se} & : & \text{F} \\ & & & \\ & \text{F} & & \end{array}$

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