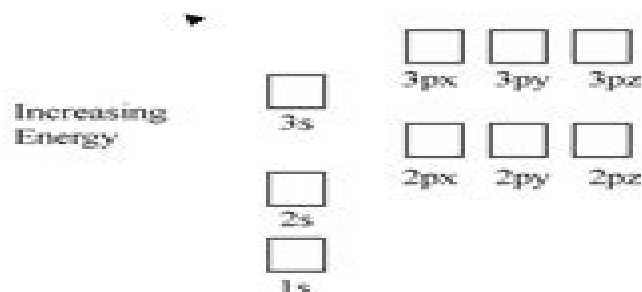


Electron configuration

The area in which there is 90% probability of finding an electron is called an orbital. An orbital can hold two electrons.

These orbitals can be represented in the following way:

Aufbau Diagram



Each box represents an orbital. Electrons are designated by arrows pointing up or down which represent clockwise or counterclockwise spins.




To determine where the electrons of a particular element are located you use the following rules:

RULE 1: An electron will always be found in the lowest energy orbital available that has room.

Therefore the electron configuration of Hydrogen is:  This is designated as: $1s^1$

RULE 2: An orbital may hold a maximum of 2 electrons.

RULE 3: The Pauli Exclusion Principle – Two electrons sharing an orbital will have opposite spins.

Therefore the electron configuration of Helium is:  This is designated as: $1s^2$

1. Draw an orbital diagram and write the electron configuration for Lithium, Beryllium, and Boron.

Li $1s^2 2s^1$

Be $1s^2 2s^2$

B $1s^2 2s^2 2p^1$

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