1)Which of the following statements is incorrect?

- A) The degree of ionization of a weak electrolyte increases with dilution.
- B) Strong electrolytes are ionized completely at all concentrations.
- C) Addition of NH<sub>4</sub>CL to NH<sub>4</sub>OH increases the ionization of the latter.
- D) Increase the temperature increases the ionization.

2) According to Lowry and Bronsted system the chloride ion (CI) in aqueous solution is a

- A) Weak base
- B) Strong base
- C) Weak acid
- D) Strong acid

In the following reactions, ZnO is respectively acting as a/an

¡)ZnO+Na<sub>2</sub>O gives Na<sub>2</sub>ZnO<sub>2</sub>

- (i) ZnO+CO, gives ZnCO,
  - A) Acid and acid
  - B) Acid and base
  - C) Base and acid
  - D) Base and base

4) Given  $k_{a1}k_{a2}$  for  $H_2S=1.1\times10^{-21}M^3$ . 1L of 0.1 M  $H_2S$  having [H\*]=0.1M contains  $S^{2^4}$  ions equal to.

- A) 6.62×103
- B) 6.62×104
- C) 6.62×10°
- D) 6.62×10°

5) Equal volume of three acid solutions of pH 3.4 and 5 are mixed in a vessel. What will be

# **Multiple Choice Questions On Chemical Equilibrium**

**Arshad Iqbal** 

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