

## SECTION A

Answer ALL questions in this Section. Write your answers in the spaces provided.

1. (a) The pH of human blood is maintained within a narrow range from 7.35 to 7.45 by a natural buffer system consisting of carbonic acid,  $\text{H}_2\text{CO}_3(\text{aq})$ , and hydrogencarbonate ions,  $\text{HCO}_3^-(\text{aq})$ .

- (i) A buffer solution containing  $\text{H}_2\text{CO}_3(\text{aq})$  and  $\text{HCO}_3^-(\text{aq})$  in equal concentrations has a pH of 6.10. Calculate the dissociation constant,  $K_a$ , for  $\text{H}_2\text{CO}_3(\text{aq})$ .

$$\text{pH} = \text{p}K_a + \log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]} \quad 1$$

In the solution,  $[\text{HCO}_3^-] = [\text{H}_2\text{CO}_3]$

$$\therefore K_a = 7.94 \times 10^{-7} \text{ mol dm}^{-3} \quad 1$$

- (ii) Calculate the ratio of concentrations of  $\text{HCO}_3^-(\text{aq})$  and  $\text{H}_2\text{CO}_3(\text{aq})$  in blood at pH 7.40.

$$7.4 = -\log(7.94 \times 10^{-7}) + \log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]} \quad 2$$

$$\log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]} = 1.3$$

$$\therefore \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]} = 20$$

(Accept answers from 19.8 to 20.0)  
(1 mark for method; 1 mark for answer)

- (iii) (f) Would the blood pH of a person increase or decrease in the course of physical exertion? Explain your answer.

During physical exertion, the concentration of  $\text{CO}_2$  in blood increases. 16

The equilibrium



shifts to the right.  $\therefore$  pH of blood will drop. 16, 16

# Marking Scheme For Chemistry Paper4 2002

**Robert R. Redfield**



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