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web the balanced chemical equation for this reaction is $\text{mg s} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg no}_3 \text{ 2 aq} + \text{h}_2 \text{ g}$ in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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and two moles of nitric acid HNO_3 react to form one mole of magnesium nitrate $\text{Mg}(\text{NO}_3)_2$ one mole of water H_2O and one mole of carbon dioxide CO_2 show chemical structure image reactants

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