

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3.
 - a. It is a number that appears in front of a formula in a chemical equation.
 - b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen $\rightarrow$ carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7.
 - a. a solution in water
 - b. a substance that accelerates a chemical reaction but can be recovered unchanged
 - c. a chemical reaction in which the products re-form the original reactants
8.
 - a. KOH
 - b. $\text{Ca}(\text{NO}_3)_2$
 - c. Na_2CO_3
 - d. CCl_4
 - e. MgBr_2
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10.
 - a. 6N
 - b. 4H, 2O
 - c. 4H, 4N, 12O
 - d. 2Ca, 4O, 4H
 - e. 3Ba, 6Cl, 18O
 - f. 5Fe, 10N, 30O
 - g. 12Mg, 8P, 32O
 - h. 4N, 16H, 2S, 8O
 - i. 12Al, 188Se, 72O
 - j. 12C, 32H
11.
 - a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
 - b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
 - c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12.
 - a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
 - b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
 - c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13.
 - a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
 - c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14.
 - a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
 - b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15.
 - a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 - b. $\text{P}_2\text{O}_5 + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
 - c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17.
 - a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 - b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
 - d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18.
 - (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$
 - (2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$
 - (3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{RX} \rightarrow \text{RY} + \text{X}$

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Table of Contents Heath Chemistry Chapter 8 Review Answers

1. Understanding the eBook Heath Chemistry Chapter 8 Review Answers
 - The Rise of Digital Reading Heath Chemistry Chapter 8 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Heath Chemistry Chapter 8 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heath Chemistry Chapter 8 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heath Chemistry Chapter 8 Review Answers
 - Personalized Recommendations
 - Heath Chemistry Chapter 8 Review Answers User Reviews and Ratings
 - Heath Chemistry Chapter 8 Review Answers and Bestseller Lists

5. Accessing Heath Chemistry Chapter 8 Review Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Heath Chemistry Chapter 8 Review Answers
 - Interactive Elements Heath Chemistry Chapter 8 Review Answers
8. Staying Engaged with Heath Chemistry Chapter 8 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heath Chemistry Chapter 8 Review Answers
9. Balancing eBooks and Physical Books Heath Chemistry Chapter 8 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heath Chemistry Chapter 8 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Heath Chemistry Chapter 8 Review Answers
 - Setting Reading Goals Heath Chemistry Chapter 8 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Heath Chemistry Chapter 8 Review Answers
 - Fact-Checking eBook Content of Heath Chemistry Chapter 8 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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